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Properties of water - Wikipedia

An acid is a molecule or ion capable of donating a proton (hydrogen ion H^+) (a Brønsted-Lowry acid), or, alternatively, capable of forming a covalent bond with an electron pair (a Lewis acid).. The first category of acids are the proton donors, or Brønsted-Lowry acids. In the special case of aqueous solutions, proton donors form the hydronium ion H_3O^+ and are known as Arrhenius acids.

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Water has a very high specific heat capacity of $4.1814 \text{ J/(g}\cdot\text{K)}$ at 25°C - the second highest among all the heteroatomic species (after ammonia), as well as a high heat of vaporization (40.65 kJ/mol or 2257 kJ/kg at the normal boiling point), both of which are a result of the extensive hydrogen bonding between its molecules. These two unusual properties allow water to moderate Earth's ...

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