

## Worksheet Chemical Equilibrium Answer Key

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A represents the “ normal ” allele and is also represented by p in the H-W equilibrium equations.  $p + q = 1$  thus,  $p = 1 - q = 1 - 0.01 = 0.99$  (c) Calculate the percentage of carriers of the trait within the population. The heterozygotes are the carriers of the trait and are represented by the  $2pq$  term in the H-W equilibrium equations.

Chemistry 12

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4.1 Writing and Balancing Chemical Equations - OpenStax  
Calculating Formal Charge from Lewis Structures Assign formal charges to each atom in the interhalogen molecule  $\text{BrCl}_3$ . Solution. Step 1. Assign one of the electrons in each  $\text{Br}-\text{Cl}$  bond to the Br atom and one to the Cl atom in that bond:

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In the first call to the function, we only define the argument  $a$ , which is a mandatory, positional argument. In the second call, we define  $a$  and  $n$ , in the order they are defined in the function. Finally, in the third call, we define  $a$  as a positional argument, and  $n$  as a keyword argument.. If all of the arguments are optional, we can even call the function with no arguments.

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In a previous unit, it was stated that all objects (regardless of their mass) free fall with the same acceleration -  $9.8 \text{ m/s}^2$ . This particular acceleration value is so important in physics that it has its own peculiar name - the acceleration of gravity - and its own peculiar symbol -  $g$ . But why do all objects free fall at the same rate of acceleration regardless of their mass?

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Balancing Equations. The chemical equation described in section 4.1 is balanced, meaning that equal numbers of atoms for each element involved in the reaction are represented on the reactant and product sides. This is a requirement the equation must satisfy to be consistent with the law of conservation of matter.

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The smaller the value of  $G^{\circ}$ , the closer the standard-state is to equilibrium. The larger the value of  $G^{\circ}$ , the further the reaction has to go to reach equilibrium. The relationship between  $G^{\circ}$  and the equilibrium constant for a chemical reaction is illustrated by the data in the table below. Values of  $G^{\circ}$  and  $K$  for Common Reactions at 25 °C

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