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5. $Ca_3(PO_4)_2 + 2H_2SO_4$ Teacher Notes on Chapter 8.
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reaction) is a key skill paramount to all other aspects of the AP Chemistry Exam. Mastery of this skill is critical, or you risk wasting valuable time on the AP Exam trying to figure out problems that should be rolling off your pen easily as if you taught a lecture on it.

The Ultimate Chemical Equations Handbook Answers Chapter 4

The Ultimate Chemical Equations Greek, or German names. Ag Sn w (ferrum) (natrium) (argentum) (stannum) (plumbum) (wolfram) iron sodium silver tin lead tungsten (cuprum) (kalium) (hydrargyrum) (stibium) (aurium) copper potassium mercury

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antimony gold Symbols are used as a sort of shorthand in writing the names of elements.

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The basic rules and table from page 13 in the Ultimate Chemical Equations Handbook. Learn with

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Answers for the questions in Chapter 8: ROUND 3
Exercise 8-1: Using the activity series, predict and balance the following single replacement reactions. Use abbreviations to indicate the appropriate phase of reactants and products where possible. Note: Not all of the reactions will occur. For those that do not, write no reaction. I.

The Ultimate Chemical Equations Handbook—Teacher Edition

Hague, G., & Smith, J. (2001). *The Ultimate Chemical Equations Handbook*. Batavia, IL: Flinn Scientific, Inc.
Correlating questions (ii) from: Rodney Schmitz (2007) Moore County Schools, NC For each of the

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following eight reactions, in part (i) write a BALANCED equation and in part (ii) answer the question about the reaction.

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Determine the oxidation number of the underlined element: KMnO_4 . Since K is an alkali metal, its charge must be $1+$. Oxygen is $2-$ but there are four of them, therefore, 4 times $2-$ equals $8-$. If $1+$ and $8-$ are added together, we get $7-$. In order for the compound to be neutral, the Mn must be $7+$.

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