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Directions: Show all work!!! No Work Shown, no credit giv ...  
If there are 10.0 g of C. 12 H 22 O 11. and 10.0 g of oxygen reacting. Which is the limiting reagent? ...  
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Identify the letter of the choice that best completes the statement or answers the question. \_\_\_\_ 1. The first step in most stoichiometry problems is to \_\_\_\_.

a. add the coefficients of the reagents  
c. convert given quantities to volumes  
b. convert given quantities to moles  
d. convert given quantities to masses \_\_\_\_ 2.

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Stoichiometry Practice Test 12.  
Using the equation in number 11,  
what is the mole ratio of carbon  
monoxide reacting to the amount  
of ... Stoichiometry is possible  
due to \_\_\_\_\_. ...  $2 \text{H}_2\text{O}$  A.  $\text{H}_2$  B.  
 $\text{O}_2$  C. Both A and B. D.  $\text{H}_2\text{O}$  .

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Answer: Aluminum bromide can  
be prepared by the reaction of  
aluminum metal with ...

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Chemistry Chapter 12: Stoichiometry. Overview of Chemistry 1 Honors Chapter 12: Stoichiometry. STUDY. PLAY. Stoichiometry. The calculation of quantities in chemical reactions is a subject of chemistry. Mole ratio. A conversion factor derived from coefficients of a balanced chemical equation interpreted in

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terms of moles.

Unit 7 : Stoichiometry - Honors  
Chemistry

e. 12 g 7. How many grams of  
nitric acid, HNO<sub>3</sub>, can be  
prepared from the reaction of 138  
g of NO<sub>2</sub> with 54.0 g H<sub>2</sub>O  
according to the equation below?

$3\text{NO}_2 + \text{H}_2\text{O} \rightarrow 2\text{HNO}_3 + \text{NO}$  ... Practice  
Test Ch 3 Stoichiometry

Name \_\_\_\_\_ Per \_\_\_\_\_  $2\text{MnO}_2 +$   
 $4\text{KOH} + \text{O}_2 + \text{Cl}_2 \rightarrow 2\text{KMnO}_4 +$   
 $2\text{KCl} + 2\text{H}_2\text{O}$  9. For the reaction  
above, there is 100. g of each  
reactant

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