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Answer In solution, the two salts ionize as indicated below. The buffering capacity of the solution is determined by the concentration ratio of proton acceptor (A^-) to proton donor (HA), or $[HPO_4^{2-}]/[H_2PO_4^-]$. From Figure 2-15, the pK_a for the dissociation of the ionizable hydrogen of dihydrogen phosphate $H_2PO_4^- \rightleftharpoons HPO_4^{2-} + H^+$ is 6.86.

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Over time, the rotation of the solution gradually decreases and reaches an equilibrium value corresponding to $[\alpha]_D^{25} = -52.5$. In contrast, a freshly prepared solution of β -D-glucose has a specific rotation of 19 . The rotation of this solution increases over time to the same equilibrium value as that shown by the α -anomer.

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