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Aqueous Solutions
Worksheet

Peptides containing
50% or more
hydrophobic residues
(W, L, I, F, M, V, Y, P,
A) are generally poorly
soluble in aqueous
solutions. We
recommend to dissolve
hydrophobic peptides in
100% organic solvent
(DMSO, DMF or

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acetonitrile) and subsequently dilute with water or buffer to the desired concentration.

Answers

Solubility Rules for Aqueous Solutions
Solubility Rules. Some combinations of aqueous reactants result in the formation of a solid precipitate as a product. However, some combinations will not

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produce such a product.

If solutions of sodium nitrate and ammonium chloride are mixed, no reaction occurs.

Peptide Solubility

Guidelines - How to

solubilize a peptide

Solubility Effects on

Reactions. Depending

on the solubility of a

solute, there are three

possible results: 1) if the

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solution has less solute than the maximum amount that it is able to dissolve (its solubility), it is a dilute solution; 2) if the amount of solute is exactly the same amount as its solubility, it is saturated; 3) if there is more solute than is able to be dissolved, the excess ...

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Chemistry LibreTexts

This table is used to

determine whether a

particular compound is

soluble or insoluble in

water (aqueous)

solution. If an insoluble

substance is formed in

the reaction between

two aqueous solutions

of different salts (ionic

compounds), it is called

a precipitate and settles

to the bottom of the

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Aqueous

Solubility - Wikipedia

Use the solubility

guidelines to predict

whether a precipitation
reaction will occur when

aqueous solutions of the
following substances are

mixed. (a) $\text{NaOH} +$

HClO_4 (b) $\text{FeCl}_2 +$

KOH (c) $(\text{NH}_4)_2\text{SO}_4$

$+ \text{NiCl}_2$

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Guidelines For

Solubility Guidelines
For Aqueous Solutions
Solubility Guidelines
for Compounds in
Aqueous Solutions.

Aqueous Solutions: An
Introduction 3)

Common nitrates,
acetates, chlorates, and
perchlorates are water
soluble.

Solubility Guidelines

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Guidelines For
for Compounds in
Aqueous Solutions
Solubility Rules for
Aqueous Solutions

□Sol.□ means that more than 3g of the substance dissolves in 100m□ of water. □Ppt.□ indicates that the combination forms a precipitate.

Solubility of Aqueous
Solutions alkali Ag, Hg,
Fe, Cu, other or NH₄ or
Pb Ba, Sr Ca Mg Zn

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metals nitrate (NO_3^-),

acetate (CH_3COO^-),

chlorate (ClO_3^-),

perchlorate (ClO_4^-)

Answers

Piersa, Amanda / Unit

10: Solutions Chemistry

Solubility of Gases in

Water - Solubility of

Ammonia, Argon,

Carbon Dioxide, Carbon

Monoxide, Chlorine,

Ethane, Ethylene,

Helium, Hydrogen,

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Guidelines For
Hydrogen Sulfide,
Methane, Nitrogen,
Oxygen and Sulfur
Dioxide in water
Solubility product

constants - The
equilibrium constant,
 K_{sp} , for aqueous
solutions of ionic
compounds at 25°C .

What Happens when
Stuff Dissolves?
The extent of the

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solubility of a substance
in a specific solvent is

measured as the

saturation concentration,

where adding more

solute does not increase

the concentration of the

solution and begins to

precipitate the excess

amount of solute.

Insolubility is the

inability to dissolve in a

solid, liquid or gaseous

solvent.

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Aqueous
Solutions, Solubility,
and Colligative
Properties ...

Answers
On the other hand,
barium phosphate
(which is likely
 $\text{BaHPO}_4(\text{s})$ in an
aqueous medium), and
silver iodide are both as
soluble as bricks. Silver
iodide precipitates from
aqueous solution as a
yellow solid, which is

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not quite so curdy (i.e. finely divided) as AgCl .

Calcium sulfate is

moderately insoluble....

Answers

Solubility guidelines for ionic compounds in water

To see all my Chemistry videos, check out <http://socratic.org/chemistry>

We'll look at what happens when you dissolve ionic and

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covalent compounds in
water. I...

Solutions

7.5: Aqueous Solutions
and Solubility -
Compounds ...

Table F: Solubility

Guidelines for Aqueous
Solutions. STUDY.

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Write. Spell. Test.

PLAY. Match. Gravity.

Created by. hwern16.

To be memorized for

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monday's optional quiz.

Includes ions, whether
they are soluble or

insoluble, and certain
exceptions if they exist.

Terms in this set (13)

group 1 ions (Li^+ , Na^+ ,
etc.)

chemistry aqueous
solutions solubility
guidelines ...

Aqueous solutions are
found throughout our

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world, and their

Aqueous

chemistry depends in

Solutions

part on how much of a

Worksheet

dissolved substance is in

Answers

them. This module

explores how substances

dissolve, why some

substances don't

dissolve, and how we

express the

concentration of a

solution. The module

describes the forces that

hold particles together

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Guidelines For

and interactions that

keep dissolved particles

apart.

Worksheet

Solved: Use the

solubility guidelines to

predict whether a ...

The number and types

of ionic charges in the

peptide determine its

solubility in aqueous

solutions. In general, the

more charged residues

the peptide possesses,

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the more soluble it is in aqueous solutions. In addition, peptides generally have more charges at pH 6–8 than at pH 2–6.

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Guidelines For

Solutions: Table F

(Solubility Guidelines

for Aqueous Solutions)

Solubility Guidelines in

Aqueous Solutions

Group 1 ions (Li^+ , Na^+ ,

etc.) Ammonium

(NH_4^+) Nitrate (NO_3^-)

Table F: Solubility

Guidelines for Aqueous

Solutions ...

Solutions: Table F

(Solubility Guidelines

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for Aqueous Solutions)

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Chemistry Table:

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Guidelines????? |

Socratic

2. What is the

concentration of an

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aqueous sodium
bromide solution with
10.7 g of solute
dissolved in 1.37 L of
water? 3. How many
grams of sodium
bromide are required to
make 13.9 mL of an
aqueous 0.15 M sodium
bromide solution? Use
the general water
solubility guidelines to
determine if the
following solutes will

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dissolve in water.
Aqueous
Solutions

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