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Shoemaker (Experiments in Physical Chemistry, 6th Ed., McGraw-Hill, Boston, 2003). Detailed

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Experimental Physical Chemistry

*Page 8/17*



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A practical - Williams College

Aim The purpose of this experiment is to determine  $x$  in the formula

$\text{Fe}(\text{NH}_4)_2(\text{SO}_4)_2 \cdot x\text{H}_2\text{O}$  by titration against a standard solution of potassium manganate (VII) (permanganate). Introduction This experiment provides useful experience in preparation for a practical examination. The directions are similar to those that would be given by an examination board.

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$5 \text{ C}_6\text{H}_5\text{COOH}(\text{s}) + 7.5 \text{ O}_2(\text{g}) \rightarrow 7 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$   
and  $n$  is equal to 7 minus 7.5, or minus 0.5.

Hence for the combustion of one mole of  
benzoic acid at 25°C  $\Delta H = -0.5 \times 8.31 \times 298 \text{ J} = -1248 \text{ J}$   
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on Fig. 1st-2. We will use benzene as the solvent. The dissolved substance (naphthalene) does not dissociate and associates in the solution. We will assume the solution is ideal, Raoult's law is

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