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Limitations of Arrhenius theory. The Arrhenius theory is applicable only in aqueous solution; for example, according to the theory, HCl is an acid in the aqueous solution but not in benzene, even though it donates H^+ ion to the benzene.

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Chemistry Chapter 7 The p

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291 C:\ChemistryXI\Unit-10\Unit-10-Lay-3(reprint).pmd
Reprint 27.7.6 THE s-BLOCK
ELEMENTS 291 The s-block
elements of the Periodic
Table are those in which the
last electron enters the
outermost s-orbital. As the s-
orbital can accommodate only
two electrons, two groups (1
& 2) belong to the s-block
of the Periodic Table.

Arrhenius theory - Arrhenius acid and base, examples

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QUESTIONS 7.1. Why are
pentahalides more covalent

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than trihalides? Ans: The group 15 elements have 5 e⁻¹s in their valence shell. It is difficult to lose 3e⁻¹s to form E³⁺ and even more difficult to lose 5e⁻¹s to form E⁵⁺. Thus, they have very little tendency to form ionic compounds.

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