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7 Chemical Formulas and Chemical Compounds

CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter SECTION 3 SHORT ANSWER

Answer the following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope.

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CHAPTER 6 REVIEW Chemical Bonding SECTION 2 SHORT ANSWER Answer the following

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questions in the space provided. 1. Use the concept of potential energy to describe how a covalent bond forms between two atoms. As the atoms involved in the formation of a covalent bond approach each other, the

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Chemical Formulas and Chemical Compounds SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In a Stock system name such as iron(III) sulfate, the Roman numeral tells us (a) how many atoms of Fe are in one formula unit. (b) how many sulfate ions can be attached to the iron atom. (c) the charge on each Fe ion.

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CHAPTER 9 REVIEW Stoichiometry MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ a. What is the value of the coefficient x in this equation? 40.07 g/mol b. What is the molar mass of C_3H_4 ? 2 mol O 2:1 mol H $2O$ c. What is the mole ratio ...

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CHAPTER 7 • HOMEWORK 7-1 VOCABULARY 1. false 2. true 3. false 4. false 5. true SKILL BUILDER $AgCl$ —In the compound, there is one atom of silver for each atom of chlorine. $Ba(OH)_2$ —The subscript 2 refers to everything inside the parentheses, so there are two hydroxide ions for

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CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter MIXED REVIEW SHORT ANSWER

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Answer the following questions in the space provided. 1. The element boron, B, has an atomic mass of 10.81 amu according to the periodic table.

6 Chemical Bonding

CHAPTER 10 REVIEW States of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Match description on the right to the correct crystal type on the left. b ionic crystal (a) has mobile electrons in the crystal c covalent molecular crystal (b) is hard, brittle, and nonconducting a metallic crystal (c) typically has the lowest melting point of the four

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HOMEWORK 7-2

the element. The modern form of the periodic table results in the (11) which states that when elements are arranged according to increasing atomic number, there is a periodic repetition of their chemical and physical (12) Study Guide for Content Mastery Chemistry: Matter and Change Chapter 6

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