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Distinguish between ionic crystals and metallic crystals. PART V Write the answers to the following questions on the line to the left, and show your work in the space provided. The molar enthalpy of fusion for water is 6.008 kJ/mol. 28.

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Part II – Problems. Solve each of the following and write your answer on the line. Be sure to include the substance and its unit. You must show all work or you will not receive any credit.
1. $N_2 + 3H_2 \rightarrow 2NH_3$ Nitrogen gas reacts with hydrogen gas to form ammonia. You have 73.5 liters of hydrogen and 35.7 liters of nitrogen gas at STP.

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SHORT ANSWER Answer the following questions in the space provided. 1. b The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products.

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Remember it is a MC test, use the answers ... 2O Then do some stoichiometry using “ easy math ” 16 g of methane (MM = 16) is 1 mole and 1 mole of methane will produce 1 mole of $CO_2 = 44$ g, and 2 moles of H_2O which is 36 g for a total of 80 g 4. d Balance: C_3H_8

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