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REVIEW Chemical
Bonding SECTION 4

SHORT ANSWER

Answer the following questions in the space provided. 1. _____ In metals, the valence electrons are considered to be (a) attached to particular positive ions. (c) immobile. (b) shared by all surrounding atoms. (d) involved in

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covalent bonds. 2.

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7 Chemical Formulas
and Chemical
Compounds

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REVIEW The Periodic
Law SECTION 1

SHORT ANSWER

Answer the following
questions in the space

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provided. 1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design. (c) according to increasing atomic number. (d) based on when they were discovered.

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Chemical Bonding In

the space provided,
write the letter of the
term or phrase that
best completes each
statement or best
answers each question.

_____ 1. The charge on
an ion is a. always
positive. b. always
negative. c. either
positive or negative. d.

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zero. _____ 2.

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CHAPTER 6

REVIEW Chemical

Bonding

SECTION 1 continued

6. Some binary
compounds are ionic,
others are covalent.

The type of bond

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avored partially
depends on the
position of the
elements in the
periodic table.

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REVIEW Reaction

Energy SECTION 2

SHORT ANSWER

Answer the following
questions in the space

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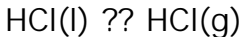
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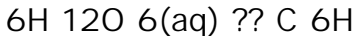
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provided. 1. For the following examples, state whether the change in entropy favors the forward or reverse reaction:

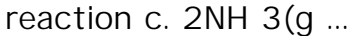
forward reaction a.



reverse reaction b. C



forward reaction c.



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Chapter 6; Chemical
Bonding Review ...

CHAPTER 6 Bonding

REVIEW Chemical

Bonding SECTION 2

SHORT ANSWER

Answer the following questions in the space provided. 1. Use the concept of potential energy to describe how a covalent bond forms between two atoms. As the atoms involved in

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