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### Modern Chemistry Chapter 3 Answers

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### 5 The Periodic Law

CHAPTER 10 REVIEW States of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Match description on the right to the correct crystal type on the left. b ionic crystal (a) has mobile electrons in the crystal c covalent molecular crystal (b) is hard, brittle, and nonconducting a metallic crystal (c) typically has the lowest melting point of the four

### Modern Chemistry Chapter 3 Homework Answers

CHAPTER 3 TEST Class Atoms: The Building Blocks of Matter ... MODERN CHEMISTRY HR W material copyçighted under notice in this work. Name CHAPTER 3 TEST continued Date Class SHORT ANSWER Write the answers to the following questions in the space provided. 22. Give three of the main concepts in Dalton's atomic theory. ...

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Holt McDougal Modern Chemistry 3 Chapter Test Chapter Test B, continued 16 Modern chemistry chapter 3 test b answers. The measure of the ability of an atom in a chemical compound to attract electrons from another atom in the compound is called \_\_\_\_\_. 17. The energy required to remove one electron from an atom is called its \_\_\_\_\_. 18.

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CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope.

### 10 States of Matter - Ms. Agostine's Chemistry Page

Holt McDougal Modern Chemistry 1 Chapter Test Assessment Chapter Test B Chapter: The Periodic Law PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. \_\_\_\_\_. 1. In his periodic table, Mendeleev did not list all of the elements in order

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### 3 Atoms: The Building Blocks of Matter

see figure 3-2 page 67. Dalton's theory and Law of Definite Proportion --see figure 3-3 page 67 . C: Modern Atomic Theory. Some of Dalton's postulates have been proven to be incorrect. His third postulate indicated that atoms cannot be subdivided. His second postulate indicates atoms of the same element have identical mass. Isotopes

### Assessment Chapter Test B

Modern Chemistry 27 . Name Section Quiz, continued Class ... CHAPTER REVIEW To give students practice under more realistic testing conditions, give them 60 minutes to answer all of the ques- tions in this Standardized Test Preparatio Answer the following items on a separate piece of paper. MULTIPLE CHOICE I. In the modern periodic table ...

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if two or more different compounds are composed of the same two elements, then the ratio of the masses of the second element combined with a certain mass of the first element is always a ratio of small whole numbers; textbook: if two or more different compounds are composed of the same two elements with a certain mass of the first element is always a ratio of small whole numbers

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### 3 Atoms: The Building Blocks of Matter - Ms. Sheehan's ...

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CHAPTER 5 REVIEW The Periodic Law SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design. (c) according to increasing atomic number. (d) based on when they were discovered. 2. d Mendeleev noticed that certain similarities in the ...

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