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141 Chapter Test

Name Class Date

Chapter Test B,

continued 18.

When a strong

acid is titrated

with a weak

base, the pH of

the solution at

the equivalence

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Answers

point is than 7.  
19. A is a highly purified solid used to check the concentration of a standard solution. 20. A 1 M solution of NaOH will have a pH that is

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**Chapters 14 & 15**

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47 Chapter Test

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Chapter Test A,

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7. The melting points of ionic compounds are higher than the melting points of molecular compounds

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because a. ionic substances tend to vaporize at room

temperature. b. ionic substances are brittle. c. attractive forces between ions are greater than the attractive ...

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The Acid-Base  
Titration and pH  
chapter of this  
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essential  
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associated with  
pH and acid-base  
titration.

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**Chapter Test B -**

**Wag & Paws**

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1 Chapter Test

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Chapter Test A

Chapter: The

Periodic Law Use

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the periodic table below to answer the questions in this Chapter Test. In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

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### **Chapter Test A -**

### **Wag & Paws**

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91 Chapter Test

Name Class Date

Chapter Test B,

continued 27.

Distinguish

between ionic

crystals and

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metallic  
crystals. PART V

Answers  
Write the

answers to the  
following  
questions on the  
line to the  
left, and show  
your work in the  
space provided.

The molar  
enthalpy of  
fusion for water  
is  $6.008 \text{ kJ/mol}$ .



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28.  
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136 Chapter Test  
Name Class Date  
Chapter Test A,  
continued

\_\_\_\_\_14. A pH  
meter measures  
the pH of a  
solution by  
measuring the a.

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Chapter 15 Test

Answers

amount of charge  
on the hydronium  
ion. b. color of  
the solution. c.  
concentration of  
the solution. d.  
voltage  
difference  
between two  
electrodes in  
the meter's  
probe. \_\_\_\_\_15.  
When you conduct  
an ...

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Chapter 16:

Reaction Energy

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Answers

by. mhardie11.  
Terms in this  
set (10) Self-  
ionization of  
water. two water  
molecules  
produce a  
hydroniumion and  
a hydroxide ion  
by transfer of a  
proton. pH.

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71 Chapter Test

Name Class Date

Chapter Test B,

continued 17. A

substance

combines with

oxygen,

releasing a

large amount of

energy as heat

and light, in

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Answers

a(n) . 18. The decomposition of a substance by an electric current is called . 19.

A(n) orders the elements by the ease with which they undergo certain chemical reactions. 20.

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## **Chapter Test A**

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I. CH 15 pH

Calculations II.

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**pH**

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Review 25

multiple choice

definition of

self-ionization

$[H_3O^+] [OH^-] = 1$

$\times 10^{-14}$

definition and

formula for

calculating  $pH =$

$-\log[H_3O^+]$   $pH > 7$

is a base;  $pH < 7$

is an acid;  $pH =$

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Chapter 15 Test

Answers

7 is neutral pH

range is

normally 0 to 14

calculate pH

from  $[H_3O^+]$

calculate  $[H_3O^+]$

from pH

definitions of

indicators,

transition ...

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Answers

What contains  
acetic acid?

Vinegar What

contains lactic  
acid? sour milk

What does

phosphoric acid

do to many

carbonated

beverages? It

gives a tart

flavor What

foods contain

citric acid?



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Chapter 15 Test

Answers

Lemons, oranges,  
grapefruits What  
contains malic  
acid? Apples  
What contains  
tartaric acid?  
Grapejuice What  
is the common  
name for Sodium  
Hydroxide NaOH?  
lye Acids were  
[...]

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3 Chapter Test

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The measure of  
the ability of  
an atom in a  
chemical  
compound to  
attract  
electrons from  
another atom in  
the compound is



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Answers

called \_\_\_\_\_.

17. The energy required to remove one electron from an atom is called its \_\_\_\_\_. 18.

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