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*Assessment Chapter Test A - Kettering City School
District*

*Modern Chemistry 47 Chapter Test Name Class Date
Chapter Test A, continued _____ 7. The melting points
of ionic compounds are higher than the melting points
of molecular compounds because a. ionic substances
tend to vaporize at room temperature. b. ionic
substances are brittle. c. attractive forces between
ions are greater than the attractive ...*

*Holt McDougal Modern Chemistry Chapter 10: States
of ...*

Modern Chemistry 84 Chapter Test Chapter: States of

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Matter In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. _____ 1. The kinetic-molecular theory of gases assumes that the particles of an ideal gas are separated by great distances. This implies that the gas

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the rates of effusion of gases at the same
temperature and pressure are inversely proportional
to the square roots of their molar masses Rate of
effusion of A / B = $\sqrt{M_b} / \sqrt{M_a}$*

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Chapter Test A, continued _____19. In the figure on the previous page, the pH at the equivalence point a. is equal to 7.0. b. is greater than 7.0. c. is less than 7.0. d. cannot be determined from the data given.
_____20. In the figure on the previous page, the volume of titration standard

Holt McDougal Modern Chemistry Chapter 11: Gases - Videos ...

Modern Chemistry 111 Chapter Test Name Class Date
Chapter Test B, continued 18. Dissolution processes with negative enthalpies of solution are processes.
19. is the solution process with water as the solvent.
20. Liquid solutes and solvents that are not soluble in

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each other are. 21. A solute molecule that is surrounded by solvent molecules is ...

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*Assessment Chapter Test A Chapter: The Periodic Law
Use the periodic table below to answer the questions
in this Chapter Test. In the space provided, write the
letter of the term or phrase that best completes each
statement or best answers each question. _____ 1.*

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*Modern Chemistry 91 Chapter Test Name Class Date
Chapter Test B, continued 27. Distinguish between
ionic crystals and metallic crystals. PART V Write the
answers to the following questions on the line to the
left, and show your work in the space provided. The
molar enthalpy of fusion for water is 6.008 kJ/mol. 28.*

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Modern Chemistry 84 Chapter Test Chapter: States of
Matter In the space provided, write the letter of the
term or phrase that best completes each statement or
best answers each question. _____ 1. The kinetic-
molecular theory of gases assumes that the particles
of an ideal gas are separated by great distances. This*

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implies that the gas

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Help ...*

*Holt McDougal Modern Chemistry 3 Chapter Test
Chapter Test B, continued 16. The measure of the
ability of an atom in a chemical compound to attract
electrons from another atom in the compound is
called _____. 17. The energy required to remove one
electron from an atom is called its _____. 18.*

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Vocabulary-New. CH 1 Mixed Questions. CH 1 Matter*

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& Energy Vocabulary-New ... CH 9 Test and CH 11 Test . CH 9 Limiting Reactant-New. CH 9 Limiting Reactant and % Yield Chapter 11 Physical Chapter 11 Molar Volume CH 11 Gases & Gas Laws.

*Assessment Chapter Test B - Wag & Paws
Modern Chemistry 71 Chapter Test Name Class Date
Chapter Test B, continued 17. A substance combines with oxygen, releasing a large amount of energy as heat and light, in a(n) . 18. The decomposition of a substance by an electric current is called . 19. A(n) orders the elements by the ease with which they undergo certain chemical reactions. 20.*

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