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3rd exam
questions

Flashcards /

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Aqueous

solutions have
pH's, and the pH
tells you (on a
logarithmic
scale) the
concentration of
H⁺ (or more
precisely H₃O⁺)
in that

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solution. A
substance is an
acid if it
lowers the pH of
a ...

Calculate the H
in a solution
that is 010 M in
NaF and 020 ...

Answer to A
solution of
Na₂SO₄ is added
dropwise to a

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*solution that is
0.010 M in Ba^{2+}
and 0.010 M in
 Sr^{2+} . (a) What
concentration...*

ChemTeam:

Molality

Problems #1-10

Calculate change

in pH based on

addition of

0.010 mol NaOH

for two

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Acidic, Basic Or
Neutral
different 1.0 L
solutions Both
solutions have
0.20 mol total
acid and

conjugate base

Solution 1:

equal amounts

(0.10 mols HA

and 0.10 mols A-

OneClass: 107. A

solution is

0.010 M Ba^{2+}

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and 0.020 M in
Acidic Basic Or

Best Answer: K_{sp}
for CuI is 1×10^{-12} . K_{sp} for
 AgI is 8.3×10^{-17} . These
two salts would
both dissociate
in the same way,
so we can say
that because AgI
has a smaller
 K_{sp} value, the

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*AgI should
precipitate
first. $[Cu^+][I^-]$
 $= 1 \times 10^{-12}$.*

*$[0.01][I^-] = 1 \times$
 10^{-12} . $[I^-] = 1$
 $\times 10^{-10}$ M.*

*$[Ag^+][I^-] = 8.3$
 $\times 10^{-17}$.*

*Solid NaI is
slowly added to
a solution that
is 0.010M in ...*

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If some solid sodium solid hydroxide is added to a solution that is 0.010-molar in $(\text{CH}_3)_3\text{CCl}$ and 0.10-molar in NaOH , which of the following is true? (Assume the temperature and volume remain

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constant.)
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Is A Solution
That 010
True/false 1.
will adding a
strong acid
(HNO₃) to a
solution that is
0.010M in
 $\text{Ag}(\text{NH}_3)_2^+$
will tend to
dissociate the

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complex ion into
 Ag^+ and NH_4^+

2. adding a
strong acid
(HNO_3) to a
solution that is
0.010M in
 $\text{AgBr}_2(^-)$ will

pH is the
Measure of
Solution Acidity
Consider a
solution that is

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*1.5 × 10⁻² M in
Ba²⁺ and*

1.9 × 10⁻² M in

*Ca²⁺. K_{sp}(BaSO₄)
= 1.07 × 10⁻¹⁰ K_{sp}(
CaSO₄) = 7.10 × 10⁻⁵*

*If sodium
sulfate is used
to selectively
precipitate one
of the cations
while leaving
the other cation
in solution,*

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*which cation
will precipitate
first? What*

minimum

*concentration of
Na₂SO₄ is*

*required to
cause the*

*precipitation of
the cation that*

*precipitates
first?*

AgNO₃ (aq) is

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*slowly added to
a solution that
is 0.280 M ...*

*If the pH of a
 6.3×10^{-3} M
solution of
codeine is 9.83
at 30 °C,
calculate the
[H⁺] of the
solution? for
the elements in
period 2 of the
periodic table*

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*describe how the
ionization
energies change?
Write one or two
sentences to
compare the
internal energy
of nitrogen gas,
N₂, with the
internal energy
of neon gas,
Ne. ?*

The pH of a

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Neutral
solution is 10
What is it's OH-
concentration

...

Consider A
Solution That Is
0.010 M In Ba^{2+}
And 0.020 M In
 Ca^{2+} . A. If
Sodium Sulfate
Is Used To
Selectively
Precipitate One
Of The Cations

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Solution That 010

Likely To Be
While Leaving
Acidic, Basic Or
The Other Cation
Neutral
In Solution,
Which Cation
Will Precipitate
First (look Up
The K_{sp} Values
In Your Text)?
What Minimum
Concentration Of
 Na_2SO_4 Is
Required To
Cause The
Precipitation Of

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Solution That 010
Likely To Be
The Cation ...
Acidic Basic Or

Neutral:
Solved:

Calculate The PH
Of A 0.010 M
CH₃CO₂H

Solution. Ka ...

"pH" = 2 The pH
of a given
solution is
nothing more
than the
negative log
base 10 of the

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Acidic, Basic Or
Neutral

concentration of
hydrogen ions,
 H^{+} , which
you'll sometimes
see written as
 H_3O^{+} ,
the hydronium
ion. . . .

Socratic Meta
Featured Answers
Topics The
[H^{+}] in a
solution is 0.01
M. What is the

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Likely To Be
pH of the
Acidic Basic Or
Neutral
Chemistry ...

*Solid NaI is
slowly added to
a solution that
is .010M in ...
pH is the
Measure of
Solution Acidity
The effective
strength of an
acid or base in*

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Likely To Be
Acidic Basic Or
Neutral

*solution, or
acidity, will
vary with both
the intrinsic
strength of the
acid or base
itself and the
amount of the
acid or base
which is present
in the aqueous
solution.*

Solved: 16.

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Consider A
Solution That Is
0.010 M In Ba^{2+}
An ...

$AgNO_3(aq)$ is
slowly added to
a solution that
is 0.280 M NaCl
and also 0.0027
M KBr? What is
 $[Ag]^+$ at the
point at which
the second anion
begins to

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precipitate?
Acidic Basic Or
Answer Save
Neutral

The $[H^+]$ in a
solution is 0.01
M. What is the
... - Socratic
(b) Proceeding
similarly for
the 0.010 M
solution, we
have Solving the
resultant
quadratic

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Acidic, Basic Or
Neutral

expression, we
obtain $x = \frac{[H^+]}{[F^-]} = 2.3 \times$

10^{-3} M . The
percentage of
molecules
ionized is
Notice that in
diluting the
solution by a
factor of 10,
the percentage
of molecules
ionized

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Acidic, Basic, Or
Neutral

Please explain!

I'm really

confused.? |

Yahoo Answers

107. A solution

is 0.010 M

Ba²⁺ and 0.020 M

in Ca²⁺. a. If

sodium sulfate

is used to

selectively

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Neutral

precipitate one
of the cations
while leaving
the other cation
in solution,
which cation
will precipitate
first? What
minimum
concentration of
 Na_2SO_4 will
trigger the
precipitation of
the cation that

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precipitates
Acidic Basic Or
first? b.
Neutral

What is the pH
of a solution
that is 0.0010 M
HCl? | Yahoo ...

Problem #1: A
solution of H₂
SO₄ with a
molal
concentration of
8.010 m has a
density of 1.354

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Neutral
g/mL. What is
the molar
concentration of
this solution?

Solution: 8.010

m means 8.010

mol / 1 kg of

solvent 8.010

mol times

98.0768 g/mol =

785.6 g of

solute

Solved: A

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Likely To Be
Solution Of
Acidic Basis Or
Neutral
*Na₂SO₄ Is Added
Dropwise To ...*

- Chegg

*The air we
breathe, what we
drink, as well
as products in
our household -
these are a few
examples of
solutions we
encounter every
day. In this*

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Solution That 010
Likely To Be
Acidic, Basic Or
Neutral

*lesson, we will
discuss what a
solution is, the
...*

*What is the H_3O^+
of a solution
with pH of 8.7 -
Answers*

*Calculate the H^+
in a solution
that is 010 M in
 NaF and 020 M in
 HF $K_a = 7.2 \times 10^{-4}$*

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from CHEM 1310
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at Georgia
Institute Of
Technology

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1989 Flashcards
| Quizlet
Actually pH is
the negative log
of the
concentration of
H⁺ ions. a
neutrl liquid

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Neutral

*water has same
H⁺ and OH⁻
concentration
which equal
(10⁻⁷) an pH =
9 solution has
hundred fold
less H⁺ ions
than a ...*

*What is a
Solution in
Science? -
Definition &*

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Examples ...

101. A solution containing sodium fluoride is mixed with one containing calcium nitrate to form a solution that is 0.015 M in NaF and 0.010 M in $\text{Ca}(\text{NO}_3)_2$. Does a precipitate form in the mixed

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Neutral

*solution? If so,
identify the
precipitate.*

*OneClass: 101. A
solution
containing
sodium fluoride
is ...*

*Answer to
Calculate the pH
of a 0.010 M
CH₃CO₂H
solution. K_a =*

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Neutral
 1.8×10^{-5} for
acetic acid.

Which equation
is NOT required
to de...

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