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3) 2, to form aluminum nitrate, $\text{Al}(\text{NO}_3)_3$? If so, write the balanced chemical equation for the reaction. Yes; because aluminum is above copper in the activity series, aluminum metal will replace copper in copper(II) nitrate. $2\text{Al}(s) + 3\text{Cu}(\text{NO}_3)_2(aq) \rightarrow 2\text{Al}(\text{NO}_3)_3(aq) + 3\text{Cu}(s)$ 70 CHEMICAL EQUATIONS AND REACTIONS MODERN CHEMISTRY

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SECTION 3 Date CHAPTER 11 REVIEW Gases Class SHORT ANSWER Answer the following questions in the space provided. c c The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (b) divided by the mass of 1 mol. nRT (c) multiplied by 22.4 L. (d) divided by 22.4 L. For the expression $V = (a) \text{ increasing } P$

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