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...

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Chemical equilibrium: A state in which the rates of the forward and reverse reactions are equal and the concentrations of the reactants and products remain constant. \Rightarrow

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which a series of measure-
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equilibrium is reached. 2.
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salt can be ground up into
small particles. The
temperature of the water can
be increased. 3. a. toward
the left b. toward the right
c. toward the right d.

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toward the right 4. a.
toward the left b. toward
the right

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change in a system at

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equilibrium. •Chemical equilibrium can be disturbed by a stress, but the system soon reaches a new equilibrium. •Le Châtelier's principle states that when a system at equilibrium is disturbed, the system

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adjusts in a way to reduce
the change. Section 3
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Equilibrium Answer the
following questions in the
space provided. 1. What is

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the difference between a reversible reaction and one that goes to completion? 2. When is chemical equilibrium reached? 3. What do the arrows in this equation indicate: $A \rightleftharpoons B$? 4. What is a complex ...

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Equilibrium 1 Reading:
Olmstead and Williams,
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sections) Purpose: The shift in equilibrium position of a chemical reaction with applied stress is determined. Introduction Chemical Equilibrium No chemical reaction goes to completion. When a reaction

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stops, some amount of
reactants remain.

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