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(ii) Among  $\text{CH}_3\text{NH}_2$  and  $(\text{C}_2\text{H}_5)_2\text{NH}$ , primarily due to the greater +I-effect of the two  $-\text{C}_2\text{H}_5$  groups over one  $-\text{CH}_3$  group,  $(\text{C}_2\text{H}_5)_2\text{NH}$  is more basic than

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$\text{C}_6\text{H}_5\text{NH}_2$ . In both  $\text{C}_6\text{H}_5\text{NH}_2$  and  $\text{C}_6\text{H}_5\text{N}(\text{CH}_3)_2$  lone pair of electrons present on N-atom is delocalized over the benzene ring but  $\text{C}_6\text{H}_5\text{N}(\text{CH}_3)_2$  is more basic due to +I effect of two- $\text{CH}_3$  groups.

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