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Objectives 3.7 The student

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*is able to identify
compounds as Brønsted-Lowry
acids, bases, and/or
conjugate acid-base pairs,
using proton-transfer
reactions to justify the
identification.*

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...

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Idea 6.*

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*1. $pH = pK_a + \log(\text{base/acid})$,
best with equimolar*

concentrations 2. $C_6H_8O_7 +$

$NaOH = NaC_6H_7O_7 + H_2O$ $C_6H_7O_7$

$+ NaOH = NaC_6H_6O_7 + H_2O$

$C_6H_6O_7 + NaOH = NaC_6H_5O_7 +$

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*H2O 3. a. Equal molar
concentrations of $C_6H_8O_7$ and
 $NaC_6H_7O_7$ b. Equal molar
concentrations of $C_6H_6O_7$ and
 $NaC_6H_5O_7$ 4. Ideal*

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Blog. 13 December 2019.

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Properties of Buffer Solutions
In this laboratory, students investigate the properties of buffer solutions. The students make two ideal buffer solutions, one consisting of a weak acid

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and its conjugate base and the other, a weak base and its conjugate acid. The initial pH of each solution is determined.

*Properties of Buffer
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*Villlanueva on ...
Equation 5 is sometimes
known as the buffer
equation; it provides the
key to calculating the
properties of buffer
solutions. When the
concentrations of the weak*

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acid and its conjugate base are equal, the ratio in Equation 5 will be equal to one and the $[H_3O^+]$ concentration will be equal to the dissociation constant K_a for the weak acid.

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Chemistry Laboratory #19.
Introduction. One of the
most important applications
of acids and bases in*

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chemistry and biology is that of buffers. A buffer solution resists rapid changes in pH when acids and bases are added to it.

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*Properties of Buffer
Solutions 241 Properties of
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AP Chemistry Review
Questions Integrating
Content, Inquiry and
Reasoning 1. The major
buffer in blood is composed*
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of the weak acid carbonic acid (H₂CO₃) and its conjugate base, bicarbonate ion (HCO₃⁻).

*Alignment for AP Chemistry
Developed in collaboration
with teacher and student*

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Unit 9: Acids and Bases Unit
10: Aqueous Equilibria*

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Idea 6 ...*

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consumer products highlights
the ability of buffers to
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*function of the respiratory
system, breathing ...*

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