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## **Classroom Resources | The Perfect Kool-Aid Concentration ...**

Apply your knowledge to calculate grams of Kool-Aid powder required to make 3 different solutions of Kool-Aid (C<sub>12</sub>H<sub>22</sub>O<sub>11</sub>) with the following concentrations: 0.2 M, 0.5 M, and 1.0 M. Determine the concentration (molarity) of properly prepared Kool-Aid through a taste test.

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## **Spectrophotometric analysis of Kool-Aid 0 - Doc Ott**

A dilution is the reduction of concentration of a solution by adding water. A serial dilution is the stepwise dilution of a substance in solution. In order to calculate the concentration of dilution, the following formula may be used: For this lab, we will assume the gram formula mass of Kool-Aid is 40. g/mol. Materials

## **Molarity and Serial Dilutions Teacher Handout**

Lab Activity- Kool-Aid Concentration

Introduction: This activity introduces you to

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solutions and allows you to experience making different concentrations of Kool-aid solution. In chemistry, concentration is usually measured by the number of moles of substance dissolved in a liter of liquid.

## **Concentration - Solutions | Saturation | Molarity - PhET ...**

You can change the concentration of the Kool Aid by increasing or decreasing the amount of Kool Aid making the drink sweeter more concentrated or less sweet. Dilution would be the process of adding more solvent to the solution. In the case of your Kool Aid,

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adding more water would decrease the concentration.

## **koolaid dilution lab**

You will be making four solutions of Kool-Aid with different concentrations of solute. You will taste the solutions to decide which one is the “correct” concentration. You will later use your data from this lab to calculate the percent concentration, by mass, of each solution and determine the best tasting concentration of Kool-Aid.

## **Kool-Aid Analysis**

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Made with Explain Everything. How to get rid of muscle knots in your neck, traps, shoulders, and back - Duration: 15:23. Tone and Tighten 2,137,339 views

### **Lab 22 - Molarity & Dilutions Lab - Google Docs**

Lab 2 Spectrophotometric Measurement of Glucose Objectives 1. Learn how to use a spectrophotometer. ... Grape Kool-Aid test solution Blank 1 Blank 2 Grape Kool-Aid Glucose ... First you will make several dilutions of a glucose standard solution that has an initial concentration of 500 mg/dl of

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glucose. See Fig. 2-3.

## **Name: Per: Lab Activity- Kool-Aid Concentration**

molarity of the solution and abbreviated as M. A dilution is taking a specific amount of a solution and adding it to a larger volume. As an example, think of making a cup of Kool-Aid. Now take  $\frac{1}{2}$  of that cup and put it in a new cup. Then add water to fill up the new cup. You have now diluted the original cup of Kool-Aid by 2!



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## **-DETERMINATION OF RED ...**

Concentration Dilution Series Kool-Aid  
kaddiamy baruh ... Making a Concentration  
Series of Solutions - Duration: 4:50. Alyssa  
... L4 - Molarity, Dilution, PPM, and  
Molality Calculations ...

## **Lab 10: What's the Concentration of Kool-Aid**

Watch your solution change color as you mix chemicals with water. Then check molarity with the concentration meter. What are all the ways you can change the concentration of your solution? Switch solutes to compare different chemicals and find out how

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concentrated you can go before you hit saturation!

### **Concentration of Kool-Aid® - ScienceGeek.net**

c230 Exp. 4 - Determination of Red#40 and Blue#1 Dyes in Kool-aid™ 3 Figure 1.

Schematic of the Serial Dilution process used for both dyes for this experiment. To prepare: 12.5 ppm dye => pipet 25 ml of 50 ppm Lab Stock Dye Solution into 100-ml volumetric flask, swirl to mix, fill to the mark and invert to mix.

**Kool-Aid Chemistry - SERC**

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Determine the concentration of the Red 40 and the Blue 1 in the solution of Kool-Aid using the standard curve data. Find the corresponding x values (concentration) for the absorbance values at the 2 l max values. Hint: use the slope of the line and y intercept from Excel.  $X = \frac{y-b}{m}$ .

### **Classroom Resources | Kool-Aid | AACT**

concentrations of Kool-aid solution. There are many ways to calculate the concentration of a substance including: molarity (M), parts per million (ppm), percent composition (% comp), and grams per liter (g/L).

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## **Lab 2 Spectrophotometric Measurement of Glucose**

Bellevue College | CHEM& 161 Page - 1 - of 9  
Beer's Law: Determining the Concentration of a Solution  
Background Color additives are used in foods for a variety of reasons. Sometimes they are used to

### **How does dilution affect concentration? + Example**

Spectrophotometric analysis of Kool-Aid → The characteristics of colored solutions have been of interest to chemists for a long time.

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Of particular interest has been the fact that colored solutions, when irradiated with white light, will selectively absorb incident light of same wavelength but not of others.

### **ANSWER KEY Lab Activity- Kool-Aid Concentration**

Solution for There is a stock solution of Kool-aid. The concentration of the stock solution is 4265.3 mg/L. A dilution is made using 0.005 L of the stock...

**Answered: There is a stock solution of Kool-aid... | bartleby**

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In this chemistry experiment, students will investigate dilution with Kool-aid. The students will use their five senses to explore the solution. Skip to Main Content  
Skip to Navigation

### **Concentration Dilution Series Kool-Aid**

Make three solutions of Kool-Aid with different concentrations and taste them to decide which is the correct concentration. You will prepare 0.1 L of each of the following Kool-Aid solutions: 0.1 M, 0.4 M, and 0.7 M. ... Lab 10: What's the Concentration of Kool-Aid ...

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## **Name Date Lab Activity- Kool-Aid Concentration**

Weigh out the calculated mass of Kool-Aid to make your 4 M solution. Place it in the plastic cup, and add enough water to dissolve the solute. Use your stir rod to mix the solution until all of the Kool-Aid is dissolved. Now, add water until the solution volume reaches the 100 mL mark.

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concentrations of Kool-aid solution. There

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