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pH and pOH. Because the constant of water, K_w is (1.0×10^{-14}) (at 25° C), the (pK_w) is 14, the constant of water determines the range of the pH scale. To understand what the pK_w is, it is important to understand first what the "p" means in pOH and pH. The addition of the "p" reflects the negative of the logarithm, $(-\log)$. Therefore, the pH is the negative logarithm of the ...

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