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Explanation: a) The energy of the photon and the electron are equal,
$$E_{\text{ph}} = E_{\text{e}} = E$$
. Their associated wavelengths are also equal,

6.2: Quantized Energy and Photons - Chemistry LibreTexts
AP CHEMISTRY

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CHAPTER REVIEW

CHAPTER 6:

ELECTRONIC

STRUCTURE AND THE

PERIODIC TABLE You

should be familiar

with the wavelike

properties of light:

frequency (ν),

wavelength (λ), and

energy (E) as well as

the equations that

show their

relationships ($E = h$

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and $c = 3 \times 10^8 \text{ m/s}$) You should be familiar with the electromagnetic spectrum and atomic spectra (bright-line spectra).

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In this section we will
discuss the energy
level of the electron
of a hydrogen atom,

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and how it changes as the electron undergoes transition. According to Bohr's theory, electrons of an atom revolve around the nucleus on certain orbits, or electron shells. Each orbit has its specific energy level, which is expressed as a negative value. This is because the electrons

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on the orbit are
"captured ...

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Answers:

Energy Level and
Transition of
Electrons | Brilliant
Math ...

In physics, an
electronvolt (symbol
eV, also written
electron-volt and
electron volt) is the
amount of kinetic

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energy gained (or lost) by a single electron accelerating from rest through an electric potential difference of one volt in vacuum. When used as a unit of energy, the numerical value of 1 eV in joules (symbol J) is equivalent to the numerical value of the charge of an

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electron in ...
Key

CHAPTER 6 -
ELECTRONIC
STRUCTURE AND THE
PERIODIC TABLE

Level 3 Physics:

Atoms – The

Photoelectric Effect -

Answers . In 2013, AS

91525 replaced AS

90522. Prior to 2013,

this was an external

standard - AS90522

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Atoms, Photons and Nuclei. It is likely to be assessed using an internal test from 2013 onwards (although teachers can select from a range of assessment techniques). There were only minor ...

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Answers and

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Key

Berühmt ...

LIGHT QUESTIONS

and ANSWERS 1.

What does it mean to say an atom is excited? An electron or electrons have more energy than their ground state, which is where they would be if they had not gained energy in some way. 2. How does the energy of

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the photon that is
emitted by an
electron jumping
between two

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Answers Along with
...

We attempted to
obtain some good
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Answers And

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Berühmt Essentials
Human Anatomy And
Physiology Answers
Galerie photo for
your needs. Here you
go. It was coming
from reliable on line
source and that we
love it. We believe it
carry interesting
things for Electron
Energy And Light
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Essentials Human
Anatomy And
Physiology Answers
...

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Electron Energy and
Light 3 11. Consider
the hydrogen
spectrum in Model 2.
a. Which color of light
has the most energy?

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b. Which color of light has the least energy? 12. Does a gas discharge tube filled with boron emit the same wavelengths of light as a tube filled with hydrogen? Use evidence from Model 2 to support your answer. 13.

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Suppose an electron
and a photon have
the same energy and

...

An electron in a
hydrogen atom

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relaxes to the $n=4$ level, emitting light of 114 THz. What is the value of n for the level in which the electron originated? Express answer as integer.

LIGHT QUESTIONS
and ANSWERS - Santa
Monica College
(ground state) to
higher energy levels

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(excited states).

Energy is released as electrons return to their lower energy levels. 18. Is energy absorbed or released for the electron transition shown in the diagram to the right? Explain.

absorbed* Electron
Energy and Light
5432

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12 Electron Energy and Light-S

If photons of light with energy less than E_0 strike a metal surface, no single photon has enough energy to eject an electron, so no electrons are emitted regardless of the intensity of the light. If a photon with energy greater than E

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o strikes the metal,
then part of its
energy is used to
overcome the forces
that hold the electron
to the ...

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the hydrogen
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b. Which color of

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light has the least energy? 12. Does a gas discharge tube filled with boron emit the same wavelengths of light as a tube filled with hydrogen? Use evidence from Model 2 to support your answer. 13.

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For an electron to move from an energy level close to the nucleus to an energy level far from the nucleus it would need to _____ energy.
(#16) answer choices

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