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Lab Math Solutions, Dilutions, Concentrations and Molarity dilutions of solutions are critical to performing accurate science. In medicine, these skills are necessary to ensuring proper medicine dosage and treatment. This kit teaches the concepts of

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molarity and dilutions using Kool-aid and nerds. By using Kool-Aid, watching the color fade in the solution provides a visual way to see

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Dilution Worksheet: 1.
How would you prepare
25.00 mL of 0.1210 M

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CH₃COOH solution
from a 0.3003 M

CH₃COOH

solution? Calculate and
fill in the plan of action.

Dilution Formula:

$$C_1V_1 = C_2V_2 \quad C_2 =$$

$$0.1210 \text{ M} \quad V_2 = 25.00$$

$$\text{mL (or 0.02500 L)} \quad C_1 =$$

$$0.3003 \text{ M} \quad V_1 = 0.1210$$

$$25.00 ;$$

CHEMISTRY

DILUTION PRACTICE

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Repeat the exercise using in an empty beaker Only 25L of spigot solution 25L Oakwood University CHEMISTRY 100 - Spring 2016 Integrated Learning Exercise 4 - Concentration and Molarity.docx . What students are saying. As a current student on this bumpy collegiate pathway, I stumbled

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4: Dilution Worksheet
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making dilutions Simple
dilutions Mixing parts
or volumes Serial
dilutions Making fixed
volumes of specific

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concentrations from
liquid reagents:

$$(C1)(V1)=(C2)(V2)$$

Percent solutions (= parts per hundred)

Molar solutions

(unit=M=moles/L)

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support under grant numbers 1246120, 1525057, and 1413739.

Laboratory Math II:
Solutions and Dilutions
Dilution Problems

Worksheet 1. How do you prepare a 250.-ml of a 2.35 M HF dilution from a 15.0 M stock solution? 2. If 455-ml of 6.0 M HNO₃ is used to make a 2.5 L dilution,

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what is the molarity of the dilution? 3. If 65.5 ml of HCl stock solution is used to make 450.-ml of a 0.675 M HCl dilution, what is

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How to Calculate
Dilution of Solutions ...
dilution and the "sub
two" refers to after
dilution. We will call
this the dilution

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equation. This equation does not have an official name like Boyle's Law, so Example #1 53.4 mL of a 1.50 M solution of NaCl is on hand, but you need some 0.800 M solution. How many mL of 0.800 M can you make? Using the dilution equation, we write:

Concentration And

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Molarity Phet Answer
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Calculations for Solutions Worksheet and Key

1) 23.5g of NaCl is dissolved in enough water to make 683L of solution.

a) What is the molarity (M) of the solution?

b) How ...

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Solutions

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Dilutions Worksheet 1)

If I add 25 mL of water to 125 mL of a 0.15 M NaOH solution, what will the molarity of the diluted solution be? 2) If I add water to 100 mL of a 0.15 M NaOH solution until the final volume is 150 mL, what will the molarity of the diluted solution be? 3) How much 0.05 M HCl solution can be made by

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diluting 250 mL of 10 M
HCl?

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329 - Everett

Community College

Dilutions Worksheet. 1)

If I have 340 mL of a 1.5
M NaBr solution, what
will the concentration

be if I add 560 mL more
water to it? 2) If I dilute
250 mL of 0.50 M

lithium acetate solution

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to a volume of 750 mL,
what will the
concentration of this
solution be?

Molarity and Serial
Dilutions Teacher
Handout

Webinar on Laboratory
Math II: Solutions and
Dilutions. This Webinar
is intended to give a
brief introduction into
the mathematics of

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making solutions commonly used in a research setting. While you may already make solutions in the lab by following recipes, we hope this Webinar will help you understand the concepts involved so that you can

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329 Everett Community College Student Support Services Program

1) If 45 mL of water are added to 250 mL of a 0.75 M K_2SO_4 solution, what will the molarity of the diluted solution be?

2) If water is added to 175 mL of a 0.45 M KOH solution until the volume is 250 mL, what ...

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1) If 45 mL ...

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Dilutions Worksheet –
Solutions 1) If I have 340 mL of a 0.5 M NaBr solution, what will the concentration be if I add 560 mL more water to it? 0.19 M (the final volume is 900 mL, set up the equation from that) 2) If I dilute 250 mL of 0.10 M lithium acetate solution to a volume of 750 mL, what will the concentration of

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Solutions

this solution be?

Molarity Review

Problems Dilutions

Worksheet

Molarity and Dilutions

Worksheet 1. Calculate

the final concentration

of a solution that is

made by dissolving 14.8

g of solid sodium

hydroxide in 600.0 mL

of solution. 2. If I add

water to 100 mL of a

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0.15 M NaOH solution until the final volume is 150 mL, what will the molarity of the diluted solution be? 3.

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Dilutions Worksheet -
Solutions 1) If I add 25 mL of water to 125 mL of a 0.15 M NaOH solution, what will the molarity of the diluted

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solution be? $M_1V_1 = M_2V_2$ (0.15 M)(125 mL) = x (150 mL) x = 0.125 M

2) If I add water to 100 mL of a 0.15 M NaOH solution until the final volume is 150 mL, what will the molarity of the diluted solution be?

$M_1V_1 = M_2V_2$...

Making Dilutions
Worksheet - Mr.
Mooney's Chemistry
Page 25/30

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Solutions
Making Dilutions

Worksheet Remember that you can change the concentration of a solution by adding more solvent. While you cannot increase the concentration of a solution in this manner, you can create a more dilute solution by increasing the amount of solvent. You can

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Solutions

determine the amount of a solution needed to dilute by using the following: $M_1 \times V_1$

concentration and molarity phet answer key - concentration ...
265 mL of solution. 2)
How many grams of beryllium chloride are needed to make 125 mL of a 0.050 M solution?
Dilutions Worksheet 1)

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Worksheet Solutions

If I add 25 mL of water to 125 mL of a 0.15 M NaOH solution, what will the molarity of the diluted solution be? 2) If I add water to 100 mL of a 0.15 M NaOH solution until the final volume is 150 mL,

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answers -

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how to calculate the dilution of solutions using this interactive quiz. Use the worksheet to identify study points to watch...

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