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View Essay - Unit 7n Nail Lab

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Reflection Assignment from SCIENCE Chemistry at Montgomery High, Skillman. Unit 7n: Nail Lab Report
Objective: The objective of this lab was to observe the amount

Stoichiometry Lab: Hard as Nails

Pre-Lab discussion for the Nail Lab.

Pre-Lab discussion for the Nail Lab.

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Iron Nails Copper Chloride

Chemistry Lab ... - Yahoo Answers

Chemical change the hypothesis was not supported due to the iron and copper sulphate chemically reacting, producing a new chemical.

Introduction: Purpose? It is hypothesized that it will be a physical

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reaction due to the iron and copper forming something new. Chemical Change...

Chemistry Nail Lab question !? | Yahoo Answers

So, we did a lab in class called Nail Lab. We had to get three iron nails and put them into copper (ii) chloride solution. They had instant rust. We measured and all. But my questions are the conclusion questions: 1) Why did the reaction stop? Which reactant was used up? How do you know? 2) Describe what was happening to the atoms of iron and copper during the reactions.

Mass nails after reaction 346 g Mass 250 ml beaker dry ...

The mass of two iron nails: At this point, the nails were put into the

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solution of copper (II) sulfate. Copper started to build up on the surface of the nail, as the iron nail reacted with the copper (II) sulfate. After the reaction was finished, the copper was scraped off the nails, collected, and washed. Mass of iron nails after reaction:

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Modeling Chemistry 1 U7 Nail Lab
2014 Name Date Pd Chemistry Unit 7
– Nail Lab Purpose The purpose of this investigation is to qualitatively and quantitatively describe the chemical reaction between copper(II) chloride and iron. Procedure Day 1 (Tues or Wed) 1. Label, then mass a plastic cup provided by the instructor. 2.

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Date Pd Chemistry Unit 7 Nail Lab

The iron atoms that just lost electrons to the copper ions then become positive iron ions and break away from the remaining iron atoms (which accounts for the thinner nails). Those positive iron atoms then drift and combine with the negative ions of chlorine that the copper ions had split from.

Introduction to Chemical Equations and Reactions Date

Nail lab – part 3 – calculations, post-lab discussion
How to make balancing equations a conceptual exercise
Overview of Ws 1 Cu-AgNO₃ lab pre-lab and part 1 Types of Reactions Lab – sample data – discuss representations and treatment of energy
PowerPoint on Ech LOLOL diagrams
Work on and w/b ws 4

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Nails in Copper (II) Chloride - Chemistry - Science Forums

Arial Times New Roman Wingdings
Franklin Gothic Demi Verdana Symbol
Century Schoolbook Helvetica New
Century Schlbk Medical design
template Unit 7n Chemical Reactions:
Particles and Energy The Nature of
Chemical Reactions Representing
chem reactions Chemical Equation
Definitions Slide 5 Other Abbreviations
Used Nail Lab Nail Lab Nail Lab Nail
...

Unit 7 Item 1: Nail Lab – Stark Science

The answer is that the system is transferring energy from Eth to Ech to produce an arrangement of particles that stores more energy. This is happening faster than energy can be

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transferred from the Eth of the surrounding molecules to the system molecules.

The Nail Lab

In chemistry we are doing the nail lab experiment. (when you add one nail in a solution of copper (II) chloride. for my moles of copper i got .033 for my moles of iron i got .031 i have to create a BCA table. Is the equation $\text{Fe} + \text{CuCl}_2 = \text{Cu} + \text{FeCl}_2$? how would you do the BCA table for this ? because they both have to be on the products side or the reactants side dont they ? i cannot put the ...

The Nail Lab

Introduction to Chemical Equations and Reactions Date_____ Purpose:
This lab will help you to see the quantitative relationship between

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reactants and products in a chemical reaction, and justify the equation that is written to describe the reaction.

Name Chemistry - Nail Lab Purpose - Mr. Kleinschrodt

An iron nail is placed in random concentrations of copper (II) chloride. Will the results be similar for each group? How much copper was produced for each mo...

Modeling Chemistry - Montgomery Township School District

We dropped iron nails in copper chloride with water solution. 1.what are any precautions you need to take? 2.explain what and aqueous solution is and give an example from the lab? 3.when you decanted identify the solution and the precipitate? 4.why did you rinse the precipitate with water

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then hydrochloric acid and again with water? 5. please write the balanced chemical equation for this ...

Unit 6 – Chemical Reactions: Particles and Energy

Stoichiometry Lab: Hard as Nails? You will consider what the coefficients of a balanced chemical equation mean in physical laboratory terms. You will react a copper(II) chloride solution with the iron in a nail.

Stoichiometry Lab Report - Weebly

5. Repeat the process for the sodium chloride, filling approx. 1/3 of the test tube. Gently tap the tube to allow the crystals to settle. 6. Push more tissue paper into the test tube on top of the white crystals. 7. Add enough water to cover the tissue paper and white crystals. 8. Obtain an iron nail and

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expose the surface by rubbing with sand ...

Modeling Chemistry Unit 7 : simplebooklet.com

Chemistry - Nail Lab Purpose The purpose is to determine the ratio of copper produced to iron consumed in a replacement reaction. Procedure Day 1 1. Label, then mass a 250 mL beaker. 2. Put 50.0 mL of copper (II) chloride in the beaker. 3. Mass 2 or 3 nails together to $\pm 0.01\text{g}$. 5. Place the nails in the copper chloride solution.

Unit 7n Nail Lab Reflection Assignment - Unit 7n Nail Lab ...

Apparently these nails have different composition, so Ferdinand's observation's cannot reliably reproduced with the nails, I have in my barn. ===== So, I took coarse

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powdered 99.9% iron (lab grade powder, not the cheap filings or powder used for magnetism experiments). With this iron powder I obtained very interesting results.

Iron and Copper Sulphate Lab by natalie clark on Prezi

The actual mass of sodium acetate at the end of the lab was 3.2 grams This measurement, when converted, is equal to 0.14 moles. The calculations I used to find this answer are below:
 $3.2\text{g} \times 1 \text{ m} \div 22.4\text{g} = 0.14$ This amount of moles is equal to 232×10 Exponent 25 molecules.

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