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change that occurs
during the solution
formation process.

solubility The
maximum amount of
solute that will dissolve

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in a given amount of solvent at a specified temperature and pressure.

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conservation of energy

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states that in any chemical reaction or physical process, energy can be converted from one form to another, but it neither created nor destroyed

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is analyzed and found

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include that seawater
is a Chemistry: Matter
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aqueous solution that

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5 5 Quantum

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CHALLENGE

PROBLEMS The state

of an electron in an

atom can be

completely described

by four quantum

numbers, designated

as n , l , m_l , and m_s . The

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first, or principal,
quantum number, n ,
indicates the
electron's
approximate distance
from the ...

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Alloys When a 58.8-g
piece of hot alloy is
placed in 125 g of cold
water in a calorimeter,
the temperature of the
alloy decreases by

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106.1C, while the temperature of the water increases by 10.5C. What is the specific heat of the alloy?

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15. Diffusion of solvent
particles across a

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semipermeable
membrane from an
area of higher solvent
concentration to an
area of lower solvent
concentration.

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To vaporize 2.00 g of
ammonia, 656 calories
are required ...

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your answers to the questions and click 'Next' to see the next set of questions. You can skip questions if you would like and come back to them later with the yellow "Go To First Skipped Question" button.

Alloys When a 58.8-g piece of hot alloy is placed in 125 g

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Energy and Chemical
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115.15.1 Objectives

Define energy.

Distinguish between potential and kinetic energy. Relate chemical potential energy to the heat lost or gained in chemical reactions. Calculate the amount of heat absorbed or released

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by a substance as its
temperature changes.

Review Vocabulary

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Figure 15.4.1 Energy
Changes
Accompanying the
Thermite Reaction

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Because enthalpy is a state function, the overall enthalpy change for the reaction of 2 mol of Al(s) with 1 mol of Fe₂O₃(s) is 851.1 kJ, whether the reaction occurs in a single step (ΔH° , shown on the left) or in three hypothetical steps (shown on the right) that involve ...

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