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ATOMIC STRUCTURE AND THE PERIODIC TABLE

CHAPTER 4 WORKSHEET PART A Given the following

isotopes, determine the atomic number, the mass number, the

number of protons, electrons and neutrons. Isotope Symbol ^{13}Li

^{31}S ^{31}U $^{\text{He}}$ $^{\text{SAI}}$ ^{81}W ^{36}Kr $^{\text{SIDU}}$ ^{37}KD Atomic Number Mass

Number Protons Electrons Neutrons Isotope Name **PART B**

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The last page of the powerpoint on Chemistry, Atoms and Ions

provides the definition of oxidation number as the charge on the

ion. You were also asked to define this term in your element

brochure. Pages 222 and 980 of your text book defines oxidation

number as the positive or negative charge of a monatomic (one

atom) ion.

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"The chemistry of an atom depends only on the number of

electrons, which equals the number of protons and is called the

atomic number. Chemistry is simply numbers, an ...

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WORKSHEET ANSWER ...

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6. 25.3% copper, 12.9% sulfur, 25.7% oxygen, 36.1% water
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