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pH and pOH The pH of a solution indicates how acidic or
basic that solution is. pH range of 0-7 acidic 7 neutral 7-14
basic Since ... $[H^+]$ pH $[OH^-]$ pOH Acidic or Basic 1. $10^{-5} M$ 5
 $10^{-9} M$ 9 Acidic 2. $10^{-7}M$ 7 $10^{-7}M$ 7 Neutral 3. $10^{-10}M$ 10
 $10^{-4} M$ 4 Basic 4. $10^{-2} M$ 2 $10^{-12}M$ 12 Acidic 5.

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aqueous (that is, water) solutions. Although this is useful
because water is a common solvent, it is limited to the
relationship between the H^+ ion and the OH^- ion. What
would be useful is a more general definition that would be
more applicable to other chemical reactions and, importantly,
independent of H_2O .

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