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Award [1] for stating that bonds require more energy to break in Na₂O than in P₄O₁₀. (ii) Phosphorus(V) oxide Sodium oxide Solid state nono Molten state noyes ;; [2] Award [2] for four correct. Award [1] for two or three correct. (c)Ammonium chloride: Accept any value in the range: 3pH 7<<; ++ + NH (aq) NH (aq) H (aq); 43 Sodium methanoate: 7pH11<<;

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1. 4.0 g of solid sodium hydroxide is added to 0.10 dm³ of 1.0 mol dm⁻³ aqueous sulfuric acid. 2NaOH (s) + H₂SO₄ (aq) → Na₂SO₄ (aq) + 2H₂O (l) Which statement is correct?

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products do not change / OWTTE; [1] (ii) 3 c 3 [glycerol] [biodiesel] =; [vegetable oil] [methanol] K × [1] (iii) to move the position of equilibrium to the right/product side / increase the yield of biodiesel; [1] (iv) no effect (on position of equilibrium); increases the rate of the forward and the reverse reactions equally

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reactants and transition state; [1] (b) gradient of the line: 63; Accept 60 to 65. E_a (= × gradient) = 0.52(kJmol⁻¹); [2] Accept 0.50 to 0.54. (c) gradient of the line less steep (less negative); [1] Accept any position as long as gradient less steep. 4. (a) loss of electrons; [1] (b) Carbon: III to IV / +3 to +4 / (+)1; Manganese:

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1.0 g of NaOH was added to 99.0 g of water. The temperature of the solution increased from 18.0 C to 20.5 C. The specific heat capacity of the solution is 4.18 J g⁻¹ K⁻¹. Give the heat evolved in kJ mol⁻¹.

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To get heat evolved = $m c \Delta T = 100 \times 4.18 \times \dots$

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