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168 Guided Reading and Study Workbook CHAPTER 16,Solutions (continued) 7. Look at Figure 16.2 on page 472. Circle the letter of each sentence that is true about a saturated solution. a. The total amount of dissolved solute remains constant. b. The total mass of undissolved crystals remains constant. c.

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Chapter 16
Chapter 16: Colligative Properties of Solutions 45 16-4. The mole fraction of (NH₄)₂SO₄(aq) is given by $x_{(\text{NH}_4)_2\text{SO}_4} = \frac{n_{(\text{NH}_4)_2\text{SO}_4}}{n_{(\text{NH}_4)_2\text{SO}_4} + n_{\text{H}_2\text{O}}}$ Because (NH₄)₂SO₄(aq) is a strong electrolyte, it dissociates completely into NH₄⁺ (aq) and SO₄²⁻ (aq) ions. Assume a one kilogram solution. The number of moles of ions in one ...

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Check your answers with those at the end of the chapter. Workbook If your instructor recommends the Active Learning Workbook, do Questions, Exercises, and Problems 17-41. Chapter 16 Assignment C: Solution Dilution and Stoichiometry Molarity is the most widely used of the four solution concentrations introduced in Assignment B.

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Chemistry Live! [Worked Solutions 2 Chapter 9 - The Mole Concept](#) 9.1 Since a sulfur atom (Ar = 32) weighs twice as much as an oxygen atom (Ar = 16), 32 g of sulfur atoms and 16 g of oxygen atoms both contain the same number of atoms.

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