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calculate the number of moles and the  
mass of a reactant or product when given  
the number of moles or the mass of

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another reactant or product. You will identify the limiting reactant in a chemical reaction. You will determine the

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Quantitative calculations involving reactions in solution are carried out in the



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same manner as we discussed in Chapter 11. Instead of masses, however, we use volumes of solutions of known concentration to determine the number of moles of reactants. Whether we are dealing with volumes of solutions of reactants or masses of reactants, the coefficients in the balanced chemical

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test b answer key. a. Two formula units  $\text{KClO}_3$  decom- pose to form two formula units  $\text{KCl}$  and three molecules  $\text{O}_2$ . b Chapter 12 stoichiometry test b answer key. Four molecules  $\text{NH}_3$  react with six molecules  $\text{NO}$  to form ?ve mol-

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