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KEY 6 Metric

System: Perform the
following metric

conversions. Show
your conversion

factors. Use correct
number of

significant figures. If
you need more

room, do

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calculations on
separate page(s).

$$0.50 \text{ m} = \underline{\quad} 5.0 \times$$

$$10^2 \underline{\quad} \text{ mm} \quad 2.00 \text{ km}$$

$$= \underline{\quad} 2.00 \times 10^3 \underline{\quad} \text{ m}$$

$$0.4000 \text{ L} = 4.000 \times$$

$$10^2 \text{ mL} \text{ or } 400.0 \text{ mL}$$

$$1.00 \text{ g} = \underline{\quad} 1.00 \times$$

$$10^{-3} \underline{\quad} \text{ kg}$$

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matter and chemical

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it may be a
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you think in ways that you have not had to in other classes before. To be successful you will need to make sure you are prepared everyday to get the most out of the class this year.

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1. matter – anything that has mass and occupies space
2. chemistry – the study of composition of matter and the changes the matter undergoes
3. organic chemistry – the study of things containing carbon
- 4.

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inorganic chemistry

– opposite of

organic chem. 5.

biochemistry – the

study of living

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2 - Chapter 1 1 a

moles of $\text{Cr}_2\text{O}_3 =$

$$20000 / 152 = 131.6$$

mol [1] moles of Al =

$$131.6 \times 2 = 263.2 \text{ mol}$$

[1] mass of Al =

$$263.2 \times 26.98 = 7100$$

g [1] b moles of

$\text{Cr}_2\text{O}_3 =$... mass of

$$\text{AgCl} = 1.50 \times 10^{-3} \times$$

$$143.32 = 0.215 \text{ g [1]}$$

c moles of $\text{AgNO}_3 =$

$$30.0 / 1000$$

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Value # of sig.
figures

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you want to study.

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going to be pictorial.

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for ...

$$1.00 \text{ mol} \times \frac{16.05 \text{ g}}{250 \text{ g}} = 0.0623$$

mol moles of $\text{CO}_2 =$

0.0623 mol volume

of $\text{CO}_2 = 0.0623 \times$

$$22.4 = 1.40 \text{ dm}^3 \text{ [1]}$$

$$16 \text{ a } PV = nRT = ; 2.00$$

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10 150 1056 8.31 300

$n \times x \times x? = x$ [1] $n =$

0.0120 mol [1] b PV

$n RT = ; 1.10 \ 10 \ 200$

1056 8.31 298 $n \times x$

$x? = x$ [1] $n = 8.88 \times$

10^{-3} mol [1] 17 a Mg

[1] b C₃H₈ [1] 18 a

HCl [1] b CO [1] 19 a

0.0100 mol [1] b 5.00

...

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Math

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Chemistry 1411 1.

See table. 2. 2 .

Valence Shell

Electron Pair

Replusion . 3. more .

This slightly

reduces the

“normal” bond

angles in the

molecular geometry.

For example, in a

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tetrahedral
geometry the normal
bond angle is 109.5° ,
but in ammonia (NH_3) - one

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