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Chapter 7: Chemical Formulas and Chemical Compounds

Chapter 7 – Chemical Formulas and Chemical Compounds 7-1 Chemical Names and Formulas I. Significance of a Chemical Formula A.

Molecular formulas 1. Number of atoms of each element in one molecule of a compound C 2H 6

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= ethane (2 carbon atoms, 6 hydrogen atoms)

B. Ionic Compounds 1.

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Chemical Formulas And Compounds Chapter SECTION 1 continued 6. Some binary compounds are ionic, others are covalent. The type of bond favored partially depends on the

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position of the elements in the periodic table.

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7 Chemical Formulas and Chemical Compounds chemists use systematic methods for naming compounds and for writing chemical formulas. In this chapter, you will be

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introduced to some of the rules used to identify simple chemical compounds.

Significance of a Chemical Formula Recall that a chemical formula indicates the relative number of atoms of each kind in a chemical compound.

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Chapter 7: Chemical Formulas and Chemical Compounds

Chapter 7 - Chemical Formulas & Chemical Compounds Chapter 7 is the final chapter of the first semester. It serves as a summary of all course content discussed so far, as the focus of this chapter is on process skills

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rather than new concepts or theory.

Chapter 7 - Chemical Formulas & Chemical Compounds - yazvac

Chapter 7 Significance of a Chemical Formula, continued • For an ionic compound – the chemical formula represents one formula unit —the simplest ratio of the cations (+) and anions (-) 2SO_4 • Note also that there is no subscript for sulfur: when there is no subscript next to an atom, the subscript is understood to be 1.

Chemical Formulas and Compounds

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Chapter 7: Chemical Formulas and Chemical Compounds Section 2: Oxidation Numbers

Oxidation numbers An oxidation number, also called oxidation state, indicates the general distribution of electrons among the bonded atoms in a molecular compound or polyatomic ion. Unlike ionic charges, oxidation numbers do not have an exact physical meaning.

CHAPTER 7 Chemical Formulas and Chemical Compounds

Importance of Chemistry Formulas Types of Chemistry Formulas or Chemical Formula How to write Chemical Formula List of Chemical

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Compound Formula. The chemical formula gives us a clear understanding regarding the chemical substances, like how many and what the chemical elements of the Periodic table consists of, along with the way they are ...

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Chemical Formula of Common Compounds, Chemical reaction ...

Chapter 7: Chemical Formulas and Chemical Compounds Section 7-1: Chemical Names and Formulas 7-1-1 Explain the significance of a chemical formula. The subscript indicates that there are 8 carbon atoms in the molecule of octane. The subscript indicates that there are 18 hydrogen atoms in the molecule of octane. Subscript 2 refers to 2 aluminum ...

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Chapter 7 – Chemical Formulas and Chemical Compounds

Chapter 6 "Chemical Bonding" Chapter 7

"Chemical Formulas and Chemical Compounds"

Chapter 8 "Writing and Balancing Equations"

Chapter 9 "Stoichiometry" Chapter 11 "Gas

Laws" Chapter 14 "Lewis Dot Structures, Acids

& Bases" VSEPR Chapters 12 & 15 "Molarity, pH

& Titrations" Chapter 18 "Chemical

Equilibrium"

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Chapter 7 - Chemical Formulas and Compounds

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Chapter 7 Section 1 Chemical Names and

Formulas Significance of a Chemical Formula,

continued •The chemical formula for an ionic

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compound represents one formula unit—the simplest ratio of the compound's positive ions (cations) and its negative ions (anions). •example: aluminum sulfate —Al₂(SO₄)₃

Chemical Formulas and Chemical Chapter 7 Compounds

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds MIXED REVIEW SHORT ANSWER

Answer the following questions in the space provided. 1. Write formulas for the following compounds: CuCO₃ a. copper(II) carbonate Na₂SO₃ b. sodium sulfite (NH₄)₃PO₄ c.

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ammonium phosphate SnS_2 d. tin(IV) sulfide
 HNO_2 e. nitrous acid 2.

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Chapter 7 Significance of a Chemical Formula

- A chemical formula indicates the relative

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number of atoms of each kind in a chemical compound. •For a molecular compound, the chemical formula reveals the number of atoms of each element contained in a single molecule of the compound. •example: octane — C_8H_{18} The subscript after the C

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