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Trending - the Periodic Table!
Honors Chem Unit 03 -
Periodicity Page 2 of 4 2.
Looking at a group, what is
the general trend in atomic
radius as you go down a
group in Model 1.

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pogil periodic trends atomic radius, is 198 pm, so the atomic radius is taken to be 99 pm or 0.99 Å. When two different atoms are bonded together their sizes can be

altered. Nonetheless, the sum of their individual atomic radii is often used as a first approximation of the bond length between two different atoms.

12.15: Periodic Trends in

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Atomic Properties - Chemistry

...

**Periodic Trends 5 Extension
Questions 15. During this
activity you may have noticed
that not all of the data
provided in the models fol-
lowed the trends. a. Identify**

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two places in Model 1 where the property listed does not fit the trend identified in this activity. b.

**15 Periodic Trends-S - MRS.
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Activities for High School
Chemistry Model 1 - Main**

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**Group Elements 1 H 2 He 37
31 1312 2372 2.1 N/A 3 Li 4
Be 5 B 6 C 7 N 8 O 9 F 10 Ne
152 112 83 77 71 66 71 70
520 900 801 1086 1402 1314
1681 ...**

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outermost electrons play important roles in determining the chemistry of an element. Knowing the trends in atomic and ionic sizes, ionization energies, and electron affinities aids in understanding chemical

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**School Chemistry Periodic
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**Unit 8 Worksheet 1 Mole
Relationships Answers,
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Worksheet, Stoichiometry
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Federalist Debate Worksheet**

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Trends- Reciprocal Style Day
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Week 10 Periodic Trends

Why? The size of an atom or ion and the attraction between the nucleus and the outermost electrons play important roles in determining the chemistry of an element.

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**Honors Chem Model 1 - Main
Group Elements Unit 03 ...
Summary of Periodic Trends.
The Periodic Table of
Elements categorizes like
elements together. Dmitri
Mendeleev, a Russian**

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scientist, was the first to create a widely accepted arrangement of the elements in 1869. Mendeleev believed that when the elements are arranged in order of increasing atomic mass, certain sets of properties

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recur periodically.

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Trends Key Questions 1.
Using only a periodic table,**

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**predict the order of
increasing atomic radius for
each of the following sets of
elements: Ca, Mg, Be Be < Mg
< Ca Ga, Br, Ge Br < Ge < Ga
Al, Tl, Si Si < Al < Tl Zn, Cd,
Fe Fe < Zn < Cd 2.**

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Explain the trend as you move
across a row of the periodic
table for each of the following
atomic properties using your**

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**understanding of effective
nuclear charge. a. Atomic
radius b. Ionization energy c.
Electronegativity Model 2 -
The Alkali Metals Atomic
Number Core Charge Atomic
Radius 1st Ionization Energy**

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