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*CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H*

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*Chapter 9: Standard Review Worksheet 1. Answers will vary. An example is included below: 2H<sub>2</sub>O<sub>2</sub> (aq) → 2H<sub>2</sub>O(l) + O<sub>2</sub> (g) This describes*

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*the decomposition reaction of hydrogen peroxide. Microscopic: Two molecules of hydrogen peroxide (in aqueous solution) decompose to produce two molecules of liquid water and one molecule of oxygen gas.*

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*9. If the percentage yield for a chemical reaction is 80.0%, and the theoretical yield is 100 grams, what is the actual yield of product?  
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Remember it is a MC test, use the answers ... 3 contains  $9 \times 10^{23}$   
oxygen atoms. III. A 200 g sample of  $\text{CaCO}_3$  contains 2 moles of  $\text{CaCO}_3$   
a. I only b. II only c. III only d. I and III only e. I, II, and III  
Practice Test Ch 3 Stoichiometry Name\_\_\_\_\_Per\_\_\_\_\_  $2\text{MnO}_2 + 4\text{KOH} + \text{O}_2$   
 $+ \text{Cl}_2 ? 2\text{KMnO}_4 + 2\text{KCl} + 2\text{H}_2\text{O}$  9. For the reaction above, there ...

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stoichiometry (which you studied in Chapter 3) deals with the mass relationships of elements in compounds. Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction. Reaction stoichiometry is the subject of this chapter and it is based on*

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