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Stoichiometry SECTION 1 SHORT ANSWER Answer the following
questions in the space provided. 1. _____ The coefficients in a
chemical equation represent the (a) masses in grams of all reactants
and products. (b) relative number of moles of reactants and products.

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CHAPTER 9 REVIEW. Stoichiometry. SECTION 9.2. PROBLEMS

Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas:

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ANSWER Answer the following questions in the space provided. 1. b

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CHAPTER 9 REVIEW Stoichiometry SECTION 9-3 PROBLEMS

Write the answer on the line to the left. Show all your work in the space provided.

1. 88% If the actual yield of a reaction is 22 g and the theoretical yield is 25 g, calculate the percent yield.

2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$ N_2 ; 2.0 mol a.

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38. To ensure that all magnesium is converted to MgO , I would use pure oxygen, not air, to carry out the reaction, because Mg could react with N_2 in air to form Mg_3N_2 . The pure oxygen should be in excess. 5.

a. 50 mol HI 6. a. 15.8

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Chapter 9 Stoichiometry Definition • Stoichiometry – Relationship between quantities • Composition stoichiometry – The mass relationships of elements in compounds (Ch 7.3) • Reaction stoichiometry – The mass relationships between reactants and products in a chemical reaction Section 1 Introduction to Stoichiometry

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Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles

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of reactants/products involved in a chemical reaction. Students had an introduction to composition stoichiometry in Chapter 3 and will now move on to some more difficult problems.

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Chapter 9 describes how to use mole ratios, molar masses, conversions, limiting reactants, and percent yield to ... Stoichiometry Review - ScienceGeek.net Homepage

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Stoichiometry. SECTION 2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(s) \rightarrow 2\text{KCl}(s) + 3\text{O}_2(g)$... CHAPTER 9 REVIEW ...

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From above we can see that if we have 12.4 mol H_2 we need 4.13 mol N_2 . We don't have that much N_2 so the .892 mol of N_2 must be the limiting reagent. We can now determine how much ammonia will be produced using the mole ratio in the balanced equation :

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CHAPTER 9 REVIEW Stoichiometry SECTION 1 SHORT

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ANSWER Answer the following questions in the space provided. 1. b
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