

## Chapter 7 Chemical Formulas And Compounds Test

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Chemistry Chapter 7 Chemical Formulas and Chemical ...

Chemistry Chapter 7 Worksheets —Chemical Formulas and Nomenclature page 1 WORKSHEET 1: Determination of oxidation number or valence number Rules to apply: 1. a. The net charges on all molecules is zero; therefore, the sum of the positive charges equals the sum of the negative charges 2. a.

CHAPTER 7 Chemical Formulas and Chemical Compounds

Chapter 7: Chemical Formulas and Chemical Compounds Section 7-1: Chemical Names and Formulas 7-1-1 Explain the significance of a chemical formula. The subscript indicates that there are 8 carbon atoms in the molecule of octane. The subscript indicates that there are 18 hydrogen atoms in the molecule of octane. Subscript 2 refers to 2 aluminum ...

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CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Label each of the following statements as True or False: True a. If the formula mass of one molecule is x amu, the molar mass is x g/mol. False b. Samples of equal numbers of moles of two different chemicals

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Chapter 7: Chemical Formulas and Chemical Compounds Section 2: Oxidation Numbers Oxidation numbers An oxidation number, also called oxidation state, indicates the general distribution of electrons among the bonded atoms in a molecular compound or polyatomic ion. Unlike ionic charges, oxidation numbers do not have an exact physical meaning.

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7. What is the empirical formula for lactic acid, which has a molecular formula of  $C_3H_5O_3$ ? a.  $CHO$  232 c.  $CHO$  363 d.  $CHO$  8.  $NaCl$  is a. a molecular formula only. b. an empirical formula only. c. both an empirical formula and a chemical formula. d. both a molecular formula and an empirical formula. 9.

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WORKSHEET 1: Determination of oxidation number or valence ...

Chapter 7 – Chemical Formulas and Chemical Compounds 7-1 Chemical Names and Formulas I. Significance of a Chemical Formula A. Molecular formulas 1. Number of atoms of each element in one molecule of a compound  $C_2H_6 =$  ethane (2 carbon atoms, 6 hydrogen atoms) B. Ionic Compounds 1.

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Chemical Formulas and Chemical Chapter 7 Compounds

Chapter 7 Significance of a Chemical Formula □A chemical formula indicates the relative number of atoms of each kind in a chemical compound. □For a molecular compound, the chemical formula reveals the number of atoms of each element contained in a single molecule of the compound. □example: octane —  $C_8H_{18}$  The subscript after the C

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chemical formula reveals the number of atoms of each element contained in a single molecule of the compound, as shown below for the hydrocarbon octane. (Hydrocarbons are molecular compounds composed solely of carbon and hydrogen.) CHEMICAL FORMULAS AND CHEMICAL COMPOUNDS 203 SECTION 7-1 OBJECTIVES Explain the significance of a chemical ...

Chapter 7: Chemical Formulas and Chemical Compounds

Chapter 7 - Chemical Formulas & Chemical Compounds Chapter 7 is the final chapter of the first semester. It serves as a summary of all course content discussed so far, as the focus of this chapter...

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CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In a Stock system name such as iron(III) sulfate, the Roman numeral tells us (a) how many atoms of Fe are in one formula unit. (b) how many sulfate ions can be attached to the iron atom. (c) the charge on ...

7 Chemical Formulas and Chemical Compounds

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. Assign the oxidation number to the specified element in each of the following examples: a. S in  $H_2SO_3$  b. S in  $MgSO_4$  c. S in  $K_2S$  d. Cu in  $Cu_2S$  e. Cr in  $Na_2CrO_4$  f. N in  $HNO_3$  g. C in  $(HCO_3)$  h. N in  $(NH_4)$  2. a.

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Chapter 7 Significance of a Chemical Formula □A chemical formula indicates the relative number of atoms of each kind in a chemical compound. □For a molecular compound, the chemical formula reveals the number of atoms of each element contained in a single molecule of the compound. □example: octane —  $C_8H_{18}$  The subscript after the C

Chapter 7 – Chemical Formulas and Chemical Compounds

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Write formulas for the following compounds:  $CuCO_3$  a. copper(II) carbonate  $Na_2SO_3$  b. sodium sulfite  $(NH_4)_3PO_4$  c. ammonium phosphate  $SnS_2$  d. tin(IV) sulfide  $HNO_2$  e. nitrous acid 2.

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