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WORKSHEET 1:
Determination of
oxidation number
or valence ...
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Compounds 7-1

Chemical Names

and Formulas I.

Significance of
a Chemical

Formula A.

Molecular
formulas 1.

Number of atoms
of each element
in one molecule

of a compound C

$2\text{H}_6 =$ ethane (2

carbon atoms, 6

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Formulas And Compounds Test ANSWER Answer

the following
questions in the
space provided.

1. Write
formulas for the
following
compounds: CuCO_3
3 a. copper(II)
carbonate Na_2SO_3
3 b. sodium
sulfite $(\text{NH}_4)_3\text{PO}_4$
4 c.

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ammonium

phosphate SnS_2

d. tin(IV)

sulfide HNO_2 e.

nitrous acid 2.

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Chapter 7

Significance of
a Chemical

Formula •A

chemical formula

indicates the

relative number

of atoms of each

kind in a

chemical

compound. •For a

molecular

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compound, the chemical formula reveals the number of atoms of each element contained in a single molecule of the compound.

•example: octane
— C₈ H₁₈ The subscript after the C

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SECTION 1 SHORT

ANSWER Answer

the following

questions in the

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space provided.

1. c In a Stock system name such as iron(III) sulfate, the Roman numeral tells us (a) how many atoms of Fe are in one formula unit.

(b) how many sulfate ions can be attached to the iron atom.

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SECTION 3 SHORT
ANSWER Answer

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Answers

the following
questions in the
space provided.

1. Label each of
the following
statements as
True or False:

True a. If the
formula mass of
one molecule is
 x amu, the molar
mass is x g/mol.

False b. Samples
of equal numbers

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SECTION 2 SHORT

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ANSWER Answer
the following

questions in the
space provided.

1. Assign the oxidation number to the specified element in each of the following examples:
a. S in H_2SO_3
b. S in MgSO_4
c. S in K_2S
d. Cu in Cu_2S
e. Cr in

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Na_2CrO_4 f. N in

HNO_3 g. C in

(HCO_3) h. N in

(NH_4) 2. a.

CHAPTER 7

Chemical

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Chapter 7:

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Section 7-1: Chemical Names and Formulas

7-1-1 Explain the significance of a chemical formula. The subscript indicates that there are 8 carbon atoms in the molecule of octane. The

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subscript
indicates that
there are 18
hydrogen atoms
in the molecule
of octane.

Subscript 2
refers to 2
aluminum ...

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Chemical
formulas.

Monatomic ions.

C and Si. Binary
ionic compound.

Show relative
number of atoms
in a chemical
compound. ions
formed from 1
atom.

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Worksheets

—Chemical

Formulas and

Nomenclature

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Formulas And Compounds Test page 1 WORKSHEET 1: Determination

of oxidation

number or

valence number

Rules to apply:

1. a. The net charges on all molecules is zero; therefore, the sum of the positive charges equals the sum of the negative

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charges 2. a.

7 Chemicals

Formulas and
Chemical
Compounds
chemical formula
reveals the
number of atoms
of each element
con-tained in a
single molecule
of the compound,
as shown below

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7 Chemical
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Answers

for the
hydrocarbon
octane.

(Hydrocarbons
are molecular
compounds com-
posed solely of
carbon and
hydrogen.)

CHEMICAL
FORMULAS AND
CHEMICAL
COMPOUNDS 203
SECTION 7-1 O

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OBJECTIVES

Explain the
significance of
a chemical ...

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final chapter of
the first

semester. It

serves as a

summary of all

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discussed so

far, as the

focus of this

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Formula •A

chemical formula

indicates the

relative number

of atoms of each

kind in a

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chemical compound. • For a molecular compound, the chemical formula reveals the number of atoms of each element contained in a single molecule of the compound. • example: octane — C_8H_{18} The subscript after

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...

Chapter 7:
Chemical
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Compounds
Section 2:
Oxidation

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Numbers

Oxidation

numbers An

oxidation

number, also

called oxidation

state, indicates

the general

distribution of

electrons among

the bonded atoms

in a molecular

compound or

polyatomic ion.

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Unlike ionic charges, oxidation numbers do not have an exact physical meaning.

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7. What is the empirical formula for lactic acid, which has a molecular formula of $C_3H_5O_3$?
- a. CH_2O
b. CHO
c. CH_2O_2
d. CHO
8. $NaCl$ is a
a. a molecular formula only.
b. an empirical

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Answers

formula only. c.

both an

empirical

formula and a

chemical

formula. d. both

a molecular

formula and an

empirical

formula. 9.

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