

# Online Library Chapter 7 Chemical Formulas And Compounds

## Chapter 7 Chemical Formulas And Compounds

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Chemistry Chapter 7 Chemical Formulas and Chemical ...

Chemistry Chapter 7 Worksheets

☐Chemical Formulas and Nomenclature

page 1 WORKSHEET 1: Determination of oxidation number or valence number

Rules to apply: 1. a. The net charges on all molecules is zero; therefore, the sum of the positive charges equals the sum of the negative charges 2. a.

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and chemical compounds, Chemical  
formulas. Monatomic ions. C and Si.

Binary ionic compound. Show relative  
number of atoms in a chemical compound.  
ions formed from 1 atom.

Chapter 7: Chemical Formulas and  
Chemical Compounds

Chapter 7 - Chemical Formulas &  
Chemical Compounds Chapter 7 is the  
final chapter of the first semester. It serves  
as a summary of all course content  
discussed so far, as the focus of this  
chapter...

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Chapter 7 Significance of a Chemical Formula  
A chemical formula indicates the relative number of atoms of each kind in a chemical compound. For a molecular compound, the chemical formula reveals the number of atoms of each element contained in a single molecule of the compound. example: octane  $C_8H_{18}$   
The subscript after the C

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Chapter 7 Chemical Formulas And  
CHAPTER 7 REVIEW Chemical  
Formulas and Chemical Compounds  
SECTION 3 SHORT ANSWER Answer  
the following questions in the space  
provided. 1. Label each of the following  
statements as True or False: True a. If the  
formula mass of one molecule is  $x$  amu,  
the molar mass is  $x$  g/mol. False b.  
Samples of equal numbers of moles of two  
different chemicals

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## 7 Chemical Formulas and Chemical Compounds

### CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds

SECTION 2 SHORT ANSWER Answer  
the following questions in the space  
provided. 1. Assign the oxidation number  
to the specified element in each of the  
following examples: a. S in  $\text{H}_2\text{SO}_3$  b. S in  
 $\text{MgSO}_4$  c. S in  $\text{K}_2\text{S}$  d. Cu in  $\text{Cu}_2\text{S}$  e. Cr in  
 $\text{Na}_2\text{CrO}_4$  f. N in  $\text{HNO}_3$  g. C in  $(\text{HCO}_3)$  h.  
N in  $(\text{NH}_4)$  2. a.

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Chemical Compounds

CHAPTER 7 REVIEW Chemical

Formulas and Chemical Compounds

MIXED REVIEW SHORT ANSWER

Answer the following questions in the space provided. 1. Write formulas for the following compounds: CuCO<sub>3</sub> a.

copper(II) carbonate Na<sub>2</sub>SO<sub>3</sub> b. sodium

sulfite (NH<sub>4</sub>)<sub>3</sub>PO<sub>4</sub> c. ammonium

phosphate SnS<sub>2</sub> d. tin(IV) sulfide HNO<sub>2</sub>

e. nitrous acid 2.

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Chemical Formulas and Chemical Chapter  
7 Compounds

Chapter 7 Significance of a Chemical  
Formula ¶A chemical formula indicates  
the relative number of atoms of each kind  
in a chemical compound. ¶For a molecular  
compound, the chemical formula reveals  
the number of atoms of each element  
contained in a single molecule of the  
compound. ¶example: octane ¶  $C_8H_{18}$   
The subscript after the C

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And Compounds Test B ...

Chapter 7: Chemical Formulas and  
Chemical Compounds Section 2:  
Oxidation Numbers Oxidation numbers  
An oxidation number, also called  
oxidation state, indicates the general  
distribution of electrons among the bonded  
atoms in a molecular compound or



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polyatomic ion. Unlike ionic charges, oxidation numbers do not have an exact physical meaning.

## Chapter 7 - Chemical Formulas & Chemical Compounds - yazvac

7. What is the empirical formula for lactic acid, which has a molecular formula of  $C_3H_5O_3$ ? a.  $CHO$  b.  $CH_2O$  c.  $CH_2O_3$  d.  $CHO_3$

8.  $NaCl$  is a. a molecular formula only. b. an empirical formula only. c. both an empirical formula and a chemical formula. d. both a molecular formula and an empirical formula.

## CHAPTER 7 Chemical Formulas and Chemical Compounds

### Chapter 7: Chemical Formulas and

### Chemical Compounds Section 7-1:

### Chemical Names and Formulas 7-1-1

Explain the significance of a chemical formula. The subscript indicates that there

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are 8 carbon atoms in the molecule of octane. The subscript indicates that there are 18 hydrogen atoms in the molecule of octane. Subscript 2 refers to 2 aluminum ...

WORKSHEET 1: Determination of oxidation number or valence ...

Chapter 7 □ Chemical Formulas and Chemical Compounds 7-1 Chemical Names and Formulas I. Significance of a Chemical Formula A. Molecular formulas 1. Number of atoms of each element in one molecule of a compound  $C_2H_6 =$  ethane (2 carbon atoms, 6 hydrogen atoms) B. Ionic Compounds 1.

7 Chemical Formulas and Chemical Compounds

chemical formula reveals the number of atoms of each element contained in a single molecule of the compound, as shown below for the hydrocarbon octane.

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(Hydrocarbons are molecular compounds composed solely of carbon and hydrogen.) CHEMICAL FORMULAS AND CHEMICAL COMPOUNDS 203 SECTION 7-1 OBJECTIVES Explain the significance of a chemical ...

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CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds  
SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In a Stock system name such as iron(III) sulfate, the Roman numeral tells us (a) how many atoms of Fe are in one formula unit. (b) how many sulfate ions can be attached to the iron atom. (c) the charge on ...

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