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Section 6.1 6.1 Ionic  
Bonding  
CHAPTER 6 REVIEW Chemical  
Bonding SECTION 6-2 SHORT  
ANSWER Answer the following  
questions in the space  
provided. 1. Use the concept  
of potential energy to

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describe how a covalent bond forms between two atoms. 2. Name two elements that form compounds that are exceptions to the octet rule.

Chapter 6 - Chemical Bonding

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Chapter 6 – Chemical Bonds

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Chapter 6 Chemical Bonding

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Ionic Bonding and Ionic  
Compounds Learn with  
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terms. mmand. Chemical  
Bonding: Chapter 6 - Section  
3. Ionic Bonding and Ionic

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Compounds. STUDY. PLAY.

Terms in this set (...)

Ionic compound. a compound composed of positive and negative ions that are combined so that the numbers of positive ...

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Chapter 6 Sections 1-5

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Chapter 6 Section 4 Section  
6-4: Metallic Bonding 6-4-3

Explain why metals are  
malleable and ductile but  
ionic-crystalline compounds

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are not. The metallic bond is the same in all directions throughout the metallic structure allowing the atoms to slide past each other. This sliding is why metals are ductile and malleable. Ionic compound

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Chapter 6: Chemical Bonding  
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Chapter 6 Chemical Bonding  
Section  
CHAPTER 6 REVIEW Chemical  
Bonding SECTION 1 SHORT  
ANSWER Answer the following

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questions in the space provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis

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structures 2. b A covalent bond consists of (a) a shared electron.

Chapter 6 Review Chemical  
Bonding Section 2&3  
Flashcards ...  
Review Previous Concepts

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Chemical Bonding CHAPTER 6  
Section 1 Introduction to  
Chemical Bonding What is a  
chemical bond and why does  
it form? Section 2 Covalent  
Bonding and Molecular  
Compounds What is a  
molecular formula? What are

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the characteristics of a  
covalent bond?

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Chemical Bonding Class Date  
Section Quiz: Metallic

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Bonding In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. 1. Chemical bonding in metals is a. the same as ionic bonding. b.

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the same as covalent bonding. c. a combination of ionic and covalent bonding.

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Modern Chemistry 47 Chemical Bonding  
CHAPTER 6 REVIEW



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Chemical Bonding SECTION 4  
SHORT ANSWER Answer the following questions in the space provided. 1. \_\_\_\_\_ In metals, the valence electrons are considered to be (a) attached to particular positive ions.

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(c) immobile. (b) shared by all surrounding atoms.(d) involved in covalent bonds.  
2.

Chapter 6 Chemical Bonds  
Section 6 2 Covalent Bonding

...

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## Chapter 6 - Chemical Bonding

In previous chapters, students have studied subatomic particles and the properties of individual atoms. In Chapter 6, we will begin studying how atoms interact with...

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relaxing after a day of activities.

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View Notes - Chapter 6

Sections 1-5 Review sheets -

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Name Date Class CHAPTER 6  
REVIEW Chemical Bonding  
SECTION 1 SHORT

Chemical Bonding: Chapter 6  
- Section 3 Flashcards |



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Quizlet

Chapter 6: Chemical Bonding.  
Section 1- Introduction to  
Chemical Bonding.

Objectives: define chemical  
bond; differentiate between  
covalent and ionic bonding;  
explain why bonding occurs;

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use the difference in electronegativity to determine whether a bond is polar covalent, nonpolar covalent or ionic ...

Section 3 IONIC BONDING AND IONIC ...

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quiz chemistry chapter 6  
chemical bonding Flashcards  
and ...

Chapter 6: Chemical Bonding,  
Modern Chemistry. This is a  
working presentation of the  
notes for this chapter.

Meaning that we may or may

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not cover all of the  
material here.

Ch 6 - HonorsChemWins  
Section 6.1 – Ionic Bonds  
When the highest occupied  
energy level of an atom is  
filled with electrons, the

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atom is stable and not likely to react. The noble gases have stable electron configurations with eight valence electrons (or two in the case of helium).

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