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6. Determine the molarity of a solution prepared by dissolving 11.85 g of KMnO<sub>4</sub> in enough water to make 750.0 mL of solution. (5 points) 7. A 0.1025g sample of copper metal is dissolved in 35.0 mL of concentrated nitric acid forming Cu<sup>2+</sup> ions in solution. The solution is then diluted to a total

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volume of 200.0 mL. Determine the molarity of the  $\text{Cu}^{2+}$ .

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