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Chemistry Chapter 3

Relative Atomic Mass

and Relative Molecular

Mass ... Chapter 3 -

Stoichiometry and

Calculations with

Formulas and

Equations: Part 3 of 5 -

Duration: 24:02.

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Answers

16. The empirical formula of garnet, a gemstone, is $\text{Fe}_3\text{Al}_2\text{Si}_3\text{O}_{12}$. An analysis of a sample of garnet gave a value of 13.8% for the mass percentage of silicon.

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3.1 One atomic mass unit is defined as a mass exactly equal to one-twelfth the mass of one carbon-12 atom. We cannot weigh a single atom, but it is possible to determine the mass of one atom relative to another experimentally.

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chemistry chemical
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Stoichiometry Chapter
3! Stoichiometry:
Calculations with
Chemical Formulas and
Equations.

Stoichiometry Anatomy
of a Chemical Equation
CH₄ (g) + 2O₂ ...

Stoichiometry Example,
one more 4NH₃ + 5O₂
2-----> 4NO + 6H₂O

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React 1.5 g of NH_3 with 2.75 g of O_2 . How much NO and H_2O is produced? What is left?

...

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3 Chemical

Stoichiometry

Stoichiometry – The

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Answers

study of quantities of materials consumed and produced in chemical reactions. ...

EXERCISE! Section 3.9 Balancing Chemical Equations Which of the following are true concerning balanced chemical

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Stoichiometry -
Michigan State
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University

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Stoichiometry: The area
of study that examines

the quantities of
substances consumed
and produced in ... of O
atoms: 1 mol H₂O, 1 mol C

O 2, 3 × 10²³
molecules O₃. Answer: 1
mol H₂O (6 × 10²³ O
atoms) < 3 × 10²³
molecules O₃ (9 × 10 ...

Practice Exercise: How
many moles of sodium

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bicarbonate are in 508g
of sodium

Answers

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Authors: Zumdahl,
Steven S., Zumdahl,

Susan A., ISBN-10:

1133611095, ISBN-13:

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Chapter 3 Practice

Problems Page 1 of 3

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The Mole Concept 1.

Calculate the mass of
 8.12×10^{22} atoms of
Mg. A. 3.28 g ...

Reaction Stoichiometry

Answer the next two
questions about the
reaction below.

Consider the following
balanced reaction: $2\text{N}_2\text{O}_5 \rightarrow 4\text{NO}_2 + \text{O}_2$

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Chapter 3 Empirical
Formulas,

Stoichiometry, and
Hydrates AP Chem
Tomei

Stoichiometry Exercises.

Answer the following to
the best of your ability.

Questions left blank are
not counted against you.

When you have
completed every
question that you desire,
click the "MARK

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TEST" button after the last exercise at the bottom of the page. A new page will appear showing your correct and incorrect responses.

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29 CHAPTER THREE
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Questions 18. The two major isotopes of boron are ^{10}B and ^{11}B . The

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listed mass of 10.81 is the average mass of a very large number of boron atoms.

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Stoichiometry 3-3 3.1a
Avogadro's Number
The mole
(abbreviated mol) is the
unit chemists use when

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counting numbers of
atoms or CHEMFILE

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CHAPTER 8

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Remember it is a MC

test, use the answers ...

Practice Test Ch3

Stoichiometry (page 3 of

3) 1. d It might be

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Answers

easiest to balance the equation with mostly

whole numbers: 2NH_3

+ $\frac{5}{2} \text{O}_2 \rightarrow 2 \text{NO}_2$

+ $3 \text{H}_2\text{O}$. The question

asks

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Avogadro ' s Number

The mole (abbreviated

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mol) is the unit chemists use when counting numbers of atoms or molecules in a sample.

The number of particles (atoms, molecules, or other objects) in one

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49 36. One method is to

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Assume each quantity of reactant is limiting, then calculate the amount of product that could be produced from each reactant. This gives two possible answers (assuming two reactants). The correct answer (the amount of product that could be produced) is always the smaller number.

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