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of electrons Reduction gain of electrons  
A redox (oxidation/reduction) process  
occurs by

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Chapter 20 Worksheet: Redox I.

Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also. ... Write half-reactions for the oxidation and reduction process for each of the following. a.  $\text{Fe}^{2+} + \text{MnO}_4^- \rightarrow \text{Fe}^{3+} + \text{Mn}^{2+}$



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+  $\text{Mn}^{2+}$  b.  $\text{Sn}^{2+} + \text{IO}_3^- \rightarrow \text{Sn}^{4+} + \text{I}^-$  c.  $\text{S}^{2-} + \text{NO}_3^- \rightarrow \text{S} + \text{NO}$  d.

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students learn the essential chemistry lessons of redox reactions.

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Chapter 19 Review Write the oxidation and reduction reactions for the following.  
Oxidation: Reduction: 2. State the

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Oxidation Number of each of the elements that is underlined. a)  $\text{NH}_3$  d)  $\text{ZnSO}_3$  b)  $\text{H}_2\text{SO}_4$  c) f)  $\text{Na}$  For the following reaction, identify the substance oxidized and the substance reduced.

Oxidation - Reduction Chapter Quiz /  
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Introduction. Oxidation-Reduction (Redox) reactions are a basic type of chemical reaction involving the transfer of electrons from one atom or chemical entity to another, or more accurately, from one type of orbital to another, that results in new bonds being formed.

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## Chapter 19: Overview of the Chapter Electrochemistry

I am going to use this Quiz to see if you fully understand the topic of Oxidation and reduction. you are gonna take this quiz after my presentation...its easy trust me. afterwards, I will grade each one of you out of 100. Good Luck! :)

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Modern Chemistry 3 Oxidation-  
Reduction Reactions CHAPTER 19  
REVIEW Oxidation-Reduction Reactions  
SECTION 1 SHORT ANSWER Answer  
the following questions in the space

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provided. 1. \_\_\_\_\_ All the following equations involve redox reactions except (a)  $\text{CaO} + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2$ . (b)  $2\text{SO}_2 + \text{O}_2 \rightarrow 2\text{SO}_3$ . (c)  $2\text{HgO} \rightarrow 2\text{Hg} + \text{O}_2$ . (d)  $\text{SnCl}_4 + 2\text{FeCl}_2 \rightarrow 2\text{FeCl}_3 + \text{SnCl}_2$ . 2.

oxidation and reduction chapter 19  
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were given in Chapter 7. These rules are summarized in Table 19-1. OXIDATION-REDUCTION REACTIONS 591

SECTION 19-1 OBJECTIVES Assign oxidation numbers to reactant and product species. Define oxidation and reduction. Explain what an oxidation-reduction reaction (redox reaction) is.

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Rule Example 1. The oxidation number of any uncombined element is 0.

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Chapter 19 Review, Section 3: Oxidation-Reduction Reactions If you need a homework assignment for your chemistry class, add this one to the possibilities.

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This assignment is composed of six questions including the identification of oxidizing and reducing agents, recognizing different redox reactions, assigning oxidation numbers, balancing equations, and more.

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Reduction Reactions

Modern Chemistry 7 Oxidation-

Reduction Reactions CHAPTER 19

REVIEW Oxidation-Reduction Reactions

SECTION 3 SHORT ANSWER Answer

the following questions in the space

provided. 1. For each of the following,

identify the stronger oxidizing or

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reducing agent. (Refer to Figure 3.2 of the text.) \_\_\_\_\_ a. Ca or Cu as a reducing agent

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Electrochemistry Review Oxidation ...

The Oxidation and Reduction Reactions  
chapter of this Holt McDougal Modern

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#### Chapter 19

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Reduction Reactions ... 6 For each of the reactions described by the following

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equations, state whether or not any oxidation and reduction is occurring, and write the oxidation-reduction half-reactions for those cases in which redox does occur. / a.  $\text{Ca(OH)}_2(\text{aq}) + 2\text{HCl}(\text{aq})$   
 $\text{CaCl}_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l}) \dots$

Glencoe Chemistry - Matter And Change



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Chapter 19: Redox ...

Section 19-1 Section 19.1 Oxidation and  
Reduction <ul><li>Describe the  
processes of oxidation and reduction.

</li></ul>spectator ion: an ion that does  
not participate in a reaction and is not  
usually shown in an ionic equation

<ul><li>Identify oxidizing and reducing

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agents.



Chapter 19: Electrochemistry Overview  
of the Chapter review oxidation-  
reduction chemistry basics galvanic cells  
spontaneous chemical reaction generates  
a voltage set-up of galvanic cell &

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identification of: anode (and half-reaction) cathode (and half-reaction) net cell reaction cell potential ( $E$  or  $E^\circ$ )

CHAPTER 19 REVIEW Oxidation-  
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Chapter 19 Review, Section 1: Oxidation-  
Reduction Reactions Here is a chemistry

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activity that was designed for a specific textbook section dealing with oxidation-reduction reactions, but it can be used in conjunction with any general chemistry curriculum.

Chapter 20 Worksheet Redox  
KEY HW Section 4 1312 S12 pt1 -

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Chapter 19 Answer Key 19.1 Strategy We follow the stepwise procedure for balancing redox reactions presented in Section. ... The oxidation half-reaction is already balanced in this regard (it contains no O atoms). The reduction half-reaction requires the addition of two H<sub>2</sub>O molecules on the product side.

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Chapter 19 Review, Section 3: Oxidation-Reduction ...

Modern Chemistry 4 Oxidation-Reduction Reactions CHAPTER 19 REVIEW Oxidation-Reduction Reactions SECTION 2 SHORT ANSWER Answer the following questions in the space

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provided. 1. \_\_\_\_\_ All of the following should be done in the process of balancing redox equations except (a) adjusting coefficients to balance atoms.

CHAPTER 19 Oxidation-Reduction  
Reactions

Chapter 19–Assignment D: Summary

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and Review The vocabulary of redox reactions and equations is sometimes confusing because the terms are so closely related. The following summary should help:

Redox Term	Change in Electrons	Change in Oxidation Number
oxidation	loss of electrons	increase
reduction	gain of electrons	decrease



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