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Bossong/Woleslagle/Tang-Johnson Chemistry Study Guide Chapter 19 Study Guide 19.1 Acid-Base Theories Define the following terms: - monoprotic, diprotic, triprotic acids - conjugate acids - conjugate bases - conjugate acid-base pairs - hydronium ion ( $\text{H}_3\text{O}^+$ ) - amphoteric - Lewis acid - Lewis base 1.

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of acid or base are added. A buffer system should contain considerably more acid than base. Specific buffer systems should be chosen based on the pH that must be maintained. 114 Study Guide for Content Mastery Chemistry: Matter and Change Chapter 19

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Study Guide: Acids and Bases 1) List at least three characteristic properties of acids and three of bases. ACIDS pH less than 7 Have a sour taste Change the color of many indicators Are corrosive (react with metals) Neutralize bases Conduct an electric current BASES

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