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Ch. 19: Acids and Bases
accept a positive hydrogen ion. Thus, an acid-base reaction in the reverse direction can occur. The acid H_3O^+ can react with the base X^- to form water and HX and the following equilibrium is established.
 $H_3O^+(aq) + X^-(aq) \rightleftharpoons H_2O(l) + HX(aq)$ Acid Base Conjugate acid Conjugate base

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Unformatted text preview: Chapter 19 "Acids, Bases, and Salts" Chemistry Section 19.1 Acid-Base Theories

OBJECTIVES: -Define the properties of acids and bases. Section 19.1 Acid-Base Theories **OBJECTIVES:**
-Compare and contrast acids and bases as defined by the theories of: a) Arrhenius, b) Brønsted-Lowry, and c) Lewis.

Chemistry - Chapter 19 Acids, Bases, and Salts - Chapter ...

Section 19.3 Strengths of Acids and Bases

- OBJECTIVES:**
 - Order bases by strength according to their base dissociation constants (K_b).

 62.

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The photo on the opposite page shows how the color of a hydrangea depends upon the presence of acids in the soil. 19.1 Acids and Bases: An Introduction 595 Figure 19-1 The blue litmus paper turned pink when it was dipped into the HCl solution, confirming that HCl is an acid.

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Chapter 19 Acids Bases Objectives

Chapter 19 Acids and Bases Objectives: Sections . End chapter problems. Due Dates. 1. Describe properties of acids and bases 2. Compare and contrast the theories of Arrhenius and Brønsted-Lowry. 19.1: 44 - 49, 72, 74, 75: 5/29: 3. Describe how H^+ and OH^- concentrations are related in an aqueous solution. 4.

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