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*phenol (aromatic alcohol) pKa~ 10*  
*alcohol pKa~ 16-18*  $O-C-H$   $C=O$   $C-C-H$   
*enol keto chemistry dominated by*  
*the keto form*  $C=O$   $H$   $sp^3$   $O-H$   
*Alcohols contain an OH group*  
*connected to a saturated carbon*  
*(sp<sup>3</sup>) Phenols contain an OH group*  
*connected to a carbon of a benzene*



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*ring 77 O H H RO R'*

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Vanderbilt University**

**Major topics: solubility product  
expressions ( $K_{sp}$ ), solubility/ $K_{sp}$   
calculations, common ion effect  
calculations, pH's effect on  
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precipitation ( $Q$  vs  $K_{sp}$ )**

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***starting structure appearing as  
shown on the graph, at which point  
would energy minimisation stop?***

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organIC ChemIstry, BloChemIstry,  
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Friday night, and you don't feel like  
cooking so you head for your  
favorite eatery, the local 1950s-style*

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*diner. There you spend an hour  
talking and laughing with*

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In this lecture I'll teach you how to  
determine and calculate how the  
common ion effect affects solubility***

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*of a saturated solution. I'll also  
show how pH cha...*

*Chapter 17 a ntroduCtIon o B  
synthetIC p*

*In basic solution,  $[OH^-] > 1 \times 10^{-7} M > [H^+]$ . Hydrogen ion cannot  
appear as a reactant because its*

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***concentration is essentially zero. If it were produced, it would instantly react with the excess hydroxide ion to produce water.***

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1 . The instantaneous rate is the***



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***rate of a reaction at any particular point in time, a period of time that is so short that the concentrations of reacta***

***Chapter 17 – Additional Aspects of Aqueous Equilibria: Part 4 of 21  
Major topics: common ion effect,***

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***definition of a buffer, pH of a buffer calculations (Henderson-Hasselbach), & predicting reactants of a buffer + acid/base.***

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