

## Chapter 15 Water And Aqueous Systems Worksheet Answers

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The high surface tension of water is due to the: a. small size of water molecules. b. low mass of water molecules. c. hydrogen bonding between water molecules. d. covalent bonds in water molecules. \_\_\_\_\_  
12. Salts and other compounds that remove moisture from air are said to be: a. efflorescent. c. colloidal.  
b. surfactant. d. hygroscopic. 10 9 ...

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Chapter 13 - States of Matter Chapter 14 - Behavior of Gases Chapter 15 - Water and Aqueous Systems  
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### **SECTION 15.1 WATER AND ITS PROPERTIES (pages 445–449)**

1. Chapter 15 “Water and Aqueous Systems” Pre-AP Chemistry Charles Page High School Stephen L. Cotton  
2. Section 15.1 Water and its Properties OBJECTIVES: –Explain the high surface tension and

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low vapor pressure of water in terms of the structure of the water molecule and hydrogen bonding. 3.

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Chapter 15 - Water and Aqueous Systems - 15.3 Heterogeneous Aqueous Systems - 15.3 Lesson Check - Page 507: 21 Answer Colloids cannot be separated through filtering because their particles are smaller than those of suspensions and they do not settle out over time.

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use these activities to yourself study the vocabulary and major concepts presented in this chapter

### **05 CTR ch15 7/12/04 8:14 AM Page 387 WATER AND AQUEOUS ...**

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to direct you.

### **Chapter 15 (Water and Aqueous Systems) Test - Chapter 15 ...**

Chapter 15 - Water and Aqueous Systems - 15.1 Water and Its Properties - 15.1 Lesson Check - Page 493: 1 Answer The hydrogen bonding in water cause the high surface tension and low vapor pressure of water.

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Chapter 15-Water and Aqueous Systems. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. Joseph\_Wang6. Terms in this set (233) Water is an amazing liquid that makes. plant and animal life on Earth Possible. In addition to water forming the Earth's oceans water, it also forms the.

### **CHAPTER 15 | Aqueous Equilibria: Chemistry of the Water World**

The Water Molecule: a Review Water is a simple tri-atomic molecule,  $H_2O$  Each O-H bond is highly polar, because of the high electronegativity of the oxygen (N, O, F, and Cl have high values) bond angle of water =  $105^\circ$  due to the bent shape, the O-H bond polarities do not cancel. This means: water is a polar molecule.

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## **Prentice Hall Chemistry Chapter 15: Water and Aqueous ...**

CHAPTER 15 | Aqueous Equilibria: Chemistry of the Water World 15.1. Collect and Organize Figure P15 shows .1 four lines to describe the possible dependence of percent ionization of acetic acid with concentration. We are to choose the one that best represents the trend for this weak acid.

## **Chapter 15 - Water and Aqueous Systems - Preston Treend**

Chapter 15 Water and Aqueous Systems 159 SECTION 15.1 WATER AND ITS PROPERTIES (pages 445–449) This section describes the properties of water in the liquid and solid states and explains how hydrogen bonding affects the surface tension and vapor pressure of water. Water in the Liquid State (pages 445–447) 1.

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## **Quia - Chapter 15 "Water and Aqueous Systems"**

Chapter 15 (Water and Aqueous Systems) Test Study Guide ? The bonds between the hydrogen and oxygen atoms in a water molecule are polar covalent bonds. ? The covalent bonds are polar because the oxygen atom has a greater electronegativity than the hydrogen atoms. ? The bonds between adjacent

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water molecules are called hydrogen bonds.

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