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**equilibrium constant for the reaction:  $2 \text{HBr(g)} \rightleftharpoons \text{H}_2 \text{(g)} + \text{Br}_2 \text{(g)}$  is  $K_c = 1.26 \times 10^{-12}$  at 500 K. This implies that: a) the product concentrations will be large relative to the reactants at equilibrium.**

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**A B; tax: a required payment to a  
local, state or national government:  
revenue: income received by a  
government from taxes and non-tax**

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**sources: tax base**

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Chapter 14: Chemical Equilibrium  
Practice Quiz 2. 1. Identify the**

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**INCORRECT statement below regarding chemical equilibrium: a) All chemical reactions are, in principle, reversible. b) Equilibrium is achieved when the forward reaction rate equals the reverse reaction rate.**

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