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## And Bases Section 3 Answers

### CHAPTER 14

### REVIEW Acids and

### Bases SECTION 1

### SHORT ANSWER

Answer the following questions in the space provided.

1. Name the following compounds as acids:

a.  $\text{H}_2\text{SO}_4$  sulfuric acid  
b.  $\text{H}_2\text{SO}_3$  sulfurous acid  
c.  $\text{H}_2\text{S}$  hydrosulfuric acid  
d.  $\text{HClO}_4$  perchloric acid  
e.  $\text{HCN}$  hydrocyanic acid



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hydrogen cyanide 2.

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Arrhenius Acid -a

hydrogen containing  
compound that ionizes  
to produce hydrogen  
ions ( H<sup>+</sup>) when  
dissolved in water

Because molecular acids  
are not made of ions,  
they cannot dissociate.



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□(aq) They must be  
pulled apart, or ionized,  
by the water. Chapter  
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in this set (51) Acids.

reacts with base to form  
salt and water. Starts  
with  $\text{h}^+$  (hydrogen ion).

Binary Acid. an acid  
that contains only two  
different elements:

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hydro- gen and one of  
the more electronegative

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ANSWER Answer the  
following questions in  
the space provided. 1. a.  
Write the two equations  
that show the two-stage  
ionization of sulfurous  
acid in water. stage 1:

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$\text{H}_2\text{SO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \sim$   
 $\text{H}_3\text{O}^+(\text{aq}) + \text{HSO}_3^-(\text{aq})$  b.

Which stage of  
ionization usually  
produces more ions?  
Explain your answer.

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AND BASES 5 When

$\text{H}_3\text{PO}_4$  is added to

water, the three acids

that are present are

$\text{H}_3\text{PO}_4$ ,  $\text{H}_2\text{PO}_4^-$ , and

$\text{HPO}_4^{2-}$ .  $\text{H}_3\text{PO}_4$ , with

the largest  $K_a$  value, is

the strongest of these

weak acids The

conjugate [DOC]

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REVIEW Acids and  
Bases MIXED

REVIEW SHORT

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ANSWER Answer the following questions in the space provided. 1. a. Write the formula for hypochlorous acid.

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### 14.1 - Acid Dissociation (Ionization) Reactions

Write the simple dissociation (ionization) reaction (omitting water) for each of the following acids: a.

Hydrochloric acid (HCl)

b. Acetic acid (HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>) c. The anilinium

ion (C<sub>6</sub>H<sub>5</sub>NH<sub>3</sub><sup>+</sup>) d.

The hydrated

aluminum(III) ion [Al(H<sub>2</sub>O)<sub>6</sub>]<sup>3+</sup>

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Acids: Definition. 1)  
sour taste, corrosive 2)  
characteristically  
changes acid-base  
indicator 3) react with  
active metals to form  
hydrogen gas 4) react  
with bases to form salt  
and water

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Chapter 14 & 15 . HW:

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SHORT ANSWER

Answer the following

questions in the space

provided. 1. Name the

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following compounds as  
acids: sulfuric acid a.  $\text{H}_2\text{SO}_4$  sulfurous acid b.  
 $\text{H}_2\text{SO}_3$  hydrosulfuric  
acid c.  $\text{H}_2\text{S}$  perchloric  
acid d.  $\text{HClO}_4$   
hydrocyanic acid e.  
hydrogen cyanide 2.

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Answers to Review #7:

Acids, Bases and Salts

Answers to Review #7:

Acids, Bases and Salts

1. Know the meanings  
of and be able to apply  
the following terms ...

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SHORT ANSWER

Answer the following questions in the space provided. 1. Name the following compounds as acids: sulfuric acid a.  $\text{H}_2\text{SO}_4$  sulfurous acid b.

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H<sub>2</sub>SO<sub>3</sub> hydrosulfuric acid  
c. H<sub>2</sub>S perchloric acid  
d. HClO<sub>4</sub>

hydrocyanic acid e.

hydrogen cyanide 2. H<sub>2</sub>S  
Which (if any) of the acids mentioned in ...

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