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SECTION 14.1 PROPERTIES OF GASES(pages 413–417)
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____ 14. The volume of a gas is doubled while the temperature is held constant. The pressure of the gas a. remains unchanged. c. is doubled. b. is reduced by one half. d. depends on the kind of gas. ____ 15. As the temperature of the gas in a balloon decreases a. the volume increases. b. the pressure increases.

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Chapter 14 The Ideal Gas Law and Its Applications
414 Chapter 14 For: Links on Gases Visit: www.SciLinks.org Web Code: cdn-1141 Kinetic theory can explain why gases are compressed more easily than liquids or solids. Gases are easily compressed because of the space between the particles in a gas. Remember that the volume of the particles in a gas is small compared to the overall volume of the gas.

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Chapter 14 The Behavior of Gases147 SECTION 14.1 PROPERTIES OF GASES(pages 413–417) This section uses kinetic theory to explain the properties of gases. This section also explains how gas pressure is affected by the amount of gas, its volume, and its temperature.

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Chapter 14– Assignment E: Volume–Volume Gas Stoichiometry Chapter 14 concludes with a section on converting between volumes of gases reacting and produced in a chemical reaction. The main idea in this section is: 1) The ratio of volumes of gases in a reaction is the same as the ratio of moles,

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Chapter 14 The Behavior of Gases. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. shirz055. Terms in this set (23) compressibility. a measure of how much the volume of matter decreases under pressure. Gases are easily compressed because. of the space between the particles in a gas.

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Chapter 14
Gas Laws Worksheet atm = 760.0 mm Hg = 101.3 kPa= 760 .0 torr Boyle ' s Law Problems: 1. If 22.5 L of nitrogen at 748 mm Hg are compressed to 725 mm Hg at constant temperature. What is the new volume? 2. A gas with a volume of 4.0L at a pressure of 205kPa is allowed to expand to a volume of 12.0L.

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426 Chapter 14 14.3 Ideal Gases Solid carbon dioxide, or dry ice, is used to protect products that need to be kept cold during shipping. The adjective dry refers to a key advantage of shipping with dry ice. Dry ice doesn ' t melt. It sublimates. Dry ice can exist because gases don ' t obey the assumptions of kinetic theory at all conditions. In

05 CTR ch14 7/12/04 8:13 AM Page 361 THE BEHAVIOR OF GASES 14
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248 CHAPTER 14 Solids, Liquids, and Gases Everywhere you go, you are surrounded by the three states of matter—solids, liquids, and gases. Look around you and find a substance in each of the three states of matter. List these substances on the lines below. ... the answers to the questions.

Chapter 14
Chapter 14 The Ideal Gas Law and Its Applications. Gases Revisited Properties of Gases (Section 4.1) Gases may be compressed. Gases expand to fill their containers uniformly. All gases have low densities compared with ... Chapter14 4th ed Cracolice super custom white.key Author:

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