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In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent: $[\text{Au}(\text{CN})_2]^-$, LaCl_3 , ethanol, and para-nitrophenol. When the limiting

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reactant is not apparent, we can determine which reactant is limiting by comparing the molar amounts of the reactants with their coefficients in the balanced chemical equation, just as we did in Section 11.4 .

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Answers 12.2 Stoichiometric

Calculations: Mole mass. These

problems are very similar, there is just one added step. Typically in a

laboratory procedure, you are

weighing grams not, moles. 1.)

Balance. 2.) Select the proper mole

ratio. 3.) Convert to moles. 4.) Convert

final answer to grams Chapter 12 ...

Chapter 12 REVIEW: Stoichiometry, Theoretical, Actual ...

By the way, about Stoichiometry

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